

Experiment 25

Calorimetry

Andy Wislock



A set of nested coffee cups is a good constant pressure calorimeter.

- To determine the specific heat of a metal
- To determine the enthalpy of neutralization for a strong acid–strong base reaction
- To determine the enthalpy of solution for the dissolution of a salt

OBJECTIVES

The following techniques are used in the Experimental Procedure:

TECHNIQUES



Accompanying all chemical and physical changes is a transfer of heat (energy); heat may be either evolved (exothermic) or absorbed (endothermic). A **calorimeter** is the laboratory apparatus that is used to measure the quantity and direction of heat flow accompanying a chemical or physical change. The heat change in chemical reactions is quantitatively expressed as the **enthalpy (or heat) of reaction**, ΔH , at constant pressure. ΔH values are negative for exothermic reactions and positive for endothermic reactions.

INTRODUCTION

ΔH values are often expressed as J/mol or kJ/mol

Three quantitative measurements of heat are detailed in this experiment: measurements of the specific heat of a metal, the heat change accompanying an acid–base reaction, and the heat change associated with the dissolution of a salt in water.

The energy (heat, expressed in joules, J) required to change the temperature of one gram of a substance by 1°C is the **specific heat**¹ of that substance:

Specific Heat of a Metal

$$\text{specific heat} \left(\frac{J}{g \cdot ^\circ\text{C}} \right) = \frac{\text{energy } (J)}{\text{mass } (g) \times \Delta T (^\circ\text{C})} \quad (25.1)$$

or, rearranging for energy,

$$\text{energy } (J) = \text{specific heat} \left(\frac{J}{g \cdot ^\circ\text{C}} \right) \times \text{mass } (g) \times \Delta T (^\circ\text{C}) \quad (25.2)$$

ΔT is the temperature change of the substance. Although the specific heat of a substance changes slightly with temperature, for our purposes, we assume it is constant over the temperature changes of this experiment.

The specific heat of a metal that does not react with water is determined by (1) heating a measured mass of the metal, M , to a known (higher) temperature; (2) placing it into a measured amount of water at a known (lower) temperature; and (3) measuring the final equilibrium temperature of the system after the two are combined.

¹The specific heat of a substance is an intensive property (independent of sample size), as are its melting point, boiling point, density, and so on.

Chemistry Principle Of Calorimetry Lab Manual

Robert Bruce Thompson



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naap esp sg naap esp sg name extrasolar planets studocu - Oct 28 2022

web complete the following sections after reviewing the background pages entitled introduction doppler shift center of mass and extrasolar planet detection question 1 label the positions on the star s orbit with the letters corresponding to the labeled positions of the radial velocity curve

extrasolar planets naap university of nebraska lincoln - May 03 2023

web the naap extrasolar planets lab introduces the search for planets outside of our solar system using the doppler and transit methods it includes simulations of the observed radial velocities of singular planetary systems and introduces the