



Chemistry Terminology

absolute temperature: This is a temperature reading made relative to absolute zero. We use the unit of Kelvins for these readings.

absolute zero: This is the lowest temperature possible. If you remember that temperature is a measurement of how much atoms move around in a solid, you can guess that they stop moving entirely at absolute zero. In reality, bonds still vibrate a little bit, but for the most part you don't see much happening.

accuracy: When you measure something, the accuracy is how close your measured value is to the real value. For example, if you're actually six feet tall and your brother measures your height as six feet, one inch, he's pretty accurate. However, if your cousin measures your height as twelve feet, 13 inches, he's not accurate at all.

acid: This is anything that gives off H^+ ions in water. Acids have a pH less than 7 and are good at dissolving metals. They turn litmus paper red and phenolphthalein colorless.

acid anhydride: This is an oxide that forms an acid when you stick it in water. An example is SO_3 - when you add water it turns into sulfuric acid, H_2SO_4 .

acid dissociation constant (Ka): This is equal to the ratio of the concentrations of an acid's conjugate base and the acid present when a weak acid dissociates in water. That is, if you have a solution of Acid X where the concentration of the conjugate base is 0.5 M and the concentration of the acid is 10 M, the acid dissociation constant is $0.5/10 = 0.05$.

activated complex: In a chemical reaction, the reagents have to join together into a great big blob before they can fall back apart into the products. This great big blob is called the activated complex (a.k.a. transition state).

activation energy: The minimum amount of energy needed for a chemical reaction to take place. For some reactions this is very small (it only takes a spark to make gasoline burn). For others, it's very high (when you burn magnesium, you need to hold it over a Bunsen burner for a minute or so).

activity series: This is when you arrange elements in the order of how much they tend to react with water and acids.

actual yield: When you do a chemical reaction, this is the amount of chemical that you actually make (i.e. The amount of stuff you can weigh).

addition reaction: A reaction where atoms add to a carbon-carbon multiple bond.

adsorption: When one substance collects on the surface of another one.

alcohol: An organic molecule containing an $-OH$ group.

aldehyde: An organic molecule containing a $-COH$ group.

alkali metals: Group I in the periodic table.

alkaline earth metals: Group II in the periodic table.

alkane: An organic molecule which contains only single carbon-carbon bonds.

alloy: A mixture of two metals. Usually, you add very small amounts of a different element to make the metal stronger and harder.

alpha particle: A radioactive particle equivalent to a helium nucleus (2 protons, 2 neutrons).

amine: An organic molecule which consists of an ammonia molecule where one or more of the hydrogen atoms has been replaced by organic groups.

amino acid: The basic building blocks of proteins. They're called "amino acids" because they're both amines (they contain nitrogen) and acids (carboxylic acids, to be precise).

amphiprotic: When something is both an acid and a base. Like amino acids, for example.

amphoteric: When something is both an acid and a base. Sounds familiar, huh?

anode: The electrode where oxidation occurs. In other words, this is where electrons are lost by a substance.

aqueous: dissolved in water.

atomic mass unit (a.m.u.): This is the smallest unit of mass we use in chemistry, and is equivalent to $1/12$ the mass of carbon-12. To all intents and purposes, protons and neutrons weigh 1 a.m.u.

atomic radius: This is one half the distance between two bonded nuclei. Why don't we just measure the distance from the nucleus to the outside of the atom - after all, isn't that the same thing as a radius? It is, but atoms are also (theoretically) infinitely large (due to quantum mechanics), making this impossible to measure.

atomic solid: A solid where there's a bunch of atoms in the lattice. This is different from an ionic solid, where ions are the things that are sticking together.

Aufbau principle: When you add protons to the nucleus to build up the elements, electrons are added into orbitals.

Avogadro's Law: If you've got two gases under the same conditions of temperature, pressure, and volume, they've got the same number of particles (atoms or molecules).

base anhydride: An oxide that forms a base when water is added. CaO is an example, turning into calcium hydroxide in water.

base: A compound that gives off OH^- ions in water. They are slippery and bitter and have a pH greater than 7.

battery: This is when a bunch of voltaic cells are stuck together.

beta particle: A radioactive particle equivalent to an electron.

bidentate ligand: A ligand that can attach twice to a metal ion.

binary compound: A compound only having two elements.

binding energy: The amount of energy that holds the neutrons and protons together in the nucleus of an atom. It's a lot of energy, which is why you don't see nuclei falling apart all over the place.



Chemistry Terminology Pamphlet Master

Michal Rosen-Zvi



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