

AP Chemistry UNIT 4 Review Exam Review Sheet

PART 1: Write/Draw the Lewis Dot Structure and the Structural formula for all of the compounds below. If the compound shows resonances, draw ALL the resonance structures for the compound. State if the Compound is an ionic or molecular compound:

1. N_2O_2 * _____

6. Beryllium chloride * _____

2. PbO_2 * _____

7. Silicon tetrabromide * _____

3. SO_3 * _____

8. Sulfur dioxide * _____

4. PCl_3 * _____

9. Carbon monoxide * _____

5. SiO_2 * _____

10. Xenon tetrafluoride * _____

PART 2: For the following compounds, use your electronegativity information on the back of your green sheet to determine the bond type (i.e. *Non-polar covalent*, *Moderately polar covalent*, *Very polar covalent*, or *Ionic*) of the compounds. For the molecular (i.e. covalent) compounds, state whether the compound is Polar or Nonpolar and why, based on the symmetry of the molecules.

1. PbO_2 4. Be_3N_2 2. PCl_3 5. SO_3 3. SiO_4 6. N_2

PART 3: Answer the following questions about bonding and intermolecular forces.

PROPERTIES	TYPE OF BOND		
	Covalent	Metallic	Ionic
Melting & Boiling Points			
Hardness			
Electrical Conductivity			
Polarity and Intermolecular			

Chemistry Unit 4 Exam

Adrian Dingle



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