

SALT ANALYSIS

SCHEME OF ANALYSIS

EXPERIMENT	OBSERVATION	INFERENCE
PRELIMINARY EXAMINATIONS :		
1) Noted the colour	Colourless	Absence of $\text{Cu}^{2+}, \text{Fe}^{2+}, \text{Fe}^{3+}, \text{Mn}^{2+}, \text{Co}^{2+}, \text{Ni}^{2+}$
	Blue	Pre. of Cu^{2+}
	Brown	Pre. of Fe^{3+}
	Green	Pre. of Ni^{2+}
	Pale pink	Pre. of Mn^{2+}
	Pink	Pre. of Co^{2+}
2) Noted the solubility a) In water b) In dil.HCl	a) Soluble and no ppt with Na_2CO_3 b) Insoluble	a) Pre. of NH_4^+ b) Pre. of Pb^{2+}
3) Flame Test : A little of the salt is made into paste with conc.HCl. A glass rod is dipped in it and introduced into the flame.	a) Blue surrounded by green b) Brick red c) Crimson red d) Bluish white e) Grassy green on prolonged heating	a) Pre. of Cu^{2+} b) Pre. of Ca^{2+} c) Pre. of Sr^{2+} d) Pre. of Pb^{2+} e) Pre. of Ba^{2+}
4) Ash Test : A little of the salt is heated with con. HNO_3 and cobalt nitrate solution. A filter paper is dipped in it and burnt.	a) Green ash b) Blue tinted ash c) Pale pink ash	a) Pre. of Zn^{2+} b) Pre. of Al^{3+} c) Pre. of Mg^{2+}
SYSTEMATIC ANALYSIS OF ANIONS :		
1) Dil. H_2SO_4 test : To the salt add dil. H_2SO_4 , warm if required.	a) Evolution of Colourless, odourless gas with brisk effervescence, turns limewater milky.	a) Pre. of carbonates
	b) Smell of vinegar	b) Pre. of acetate

Cbse Salt Analysis Manual

VM Jensen



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Table of Contents Cbse Salt Analysis Manual

1. Understanding the eBook Cbse Salt Analysis Manual
 - The Rise of Digital Reading Cbse Salt Analysis Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Cbse Salt Analysis Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Cbse Salt Analysis Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Cbse Salt Analysis Manual
 - Personalized Recommendations
 - Cbse Salt Analysis Manual User Reviews and Ratings
 - Cbse Salt Analysis Manual and Bestseller Lists
5. Accessing Cbse Salt Analysis Manual Free and Paid eBooks
 - Cbse Salt Analysis Manual Public Domain eBooks
 - Cbse Salt Analysis Manual eBook Subscription Services
 - Cbse Salt Analysis Manual Budget-Friendly Options
6. Navigating Cbse Salt Analysis Manual eBook Formats

- ePub, PDF, MOBI, and More
- Cbse Salt Analysis Manual Compatibility with Devices
- Cbse Salt Analysis Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Cbse Salt Analysis Manual
 - Highlighting and Note-Taking Cbse Salt Analysis Manual
 - Interactive Elements Cbse Salt Analysis Manual
- 8. Staying Engaged with Cbse Salt Analysis Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Cbse Salt Analysis Manual
- 9. Balancing eBooks and Physical Books Cbse Salt Analysis Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Cbse Salt Analysis Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Cbse Salt Analysis Manual
 - Setting Reading Goals Cbse Salt Analysis Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Cbse Salt Analysis Manual
 - Fact-Checking eBook Content of Cbse Salt Analysis Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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