

AP Chemistry UNIT 4 Review Exam Review Sheet

PART 1: Write/Draw the Lewis Dot Structure and the Structural formula for all of the compounds below. If the compound shows resonances, draw ALL the resonance structures for the compound. State if the Compound is an ionic or molecular compound:

1. N_2O_2 * _____

6. Beryllium chloride * _____

2. PbO_2 * _____

7. Silicon tetrabromide * _____

3. SO_3 * _____

8. Sulfur dioxide * _____

4. PCl_3 * _____

9. Carbon monoxide * _____

5. SiO_2 * _____

10. Xenon tetrafluoride * _____

PART 2: For the following compounds, use your electronegativity information on the back of your green sheet to determine the bond type (i.e. *Non-polar covalent*, *Moderately polar covalent*, *Very polar covalent*, or *Ionic*) of the compounds. For the molecular (i.e. covalent) compounds, state whether the compound is Polar or Nonpolar and why, based on the symmetry of the molecules.

1. PbO_2 4. Be_3N_2 2. PCl_3 5. SO_3 3. SiO_4 6. N_2

PART 3: Answer the following questions about bonding and intermolecular forces.

PROPERTIES	TYPE OF BOND		
	Covalent	Metallic	Ionic
Melting & Boiling Points			
Hardness			
Electrical Conductivity			
Polarity and Intermolecular			

Chemistry Unit 4 Exam

S Baum

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'S Baum' text.

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Table of Contents Chemistry Unit 4 Exam

1. Understanding the eBook Chemistry Unit 4 Exam
 - The Rise of Digital Reading Chemistry Unit 4 Exam
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 4 Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Unit 4 Exam
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 4 Exam
 - Personalized Recommendations
 - Chemistry Unit 4 Exam User Reviews and Ratings

- Chemistry Unit 4 Exam and Bestseller Lists
- 5. Accessing Chemistry Unit 4 Exam Free and Paid eBooks
 - Chemistry Unit 4 Exam Public Domain eBooks
 - Chemistry Unit 4 Exam eBook Subscription Services
 - Chemistry Unit 4 Exam Budget-Friendly Options
- 6. Navigating Chemistry Unit 4 Exam eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Unit 4 Exam Compatibility with Devices
 - Chemistry Unit 4 Exam Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Unit 4 Exam
 - Highlighting and Note-Taking Chemistry Unit 4 Exam
 - Interactive Elements Chemistry Unit 4 Exam
- 8. Staying Engaged with Chemistry Unit 4 Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Unit 4 Exam
- 9. Balancing eBooks and Physical Books Chemistry Unit 4 Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Unit 4 Exam
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Unit 4 Exam
 - Setting Reading Goals Chemistry Unit 4 Exam
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Unit 4 Exam
 - Fact-Checking eBook Content of Chemistry Unit 4 Exam
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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