

Periodic Trends

1. Explain the following trends:

- Atomic Radius
- Electronegativity
- Ionization Energy

2. Order the following elements by increasing atomic radius:

Li, Na, K, Rb, Cs, Fr

3. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

4. An element has the following properties:

- Atomic Number: 19
- Electronegativity: 0.9
- Ionization Energy: 418 kJ/mol

5. Order the following elements by increasing ionization energy:

Li, Na, K, Rb, Cs, Fr

6. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

7. Order the following elements by increasing atomic radius:

Li, Na, K, Rb, Cs, Fr

8. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

9. Order the following elements by increasing ionization energy:

Li, Na, K, Rb, Cs, Fr

10. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

Name: _____ Date: _____ Period: _____

Periodic Trends Test Study Guide

Atomic Structure and Properties of the Periodic Table

1. What are the three subatomic particles found in the periodic table?

- Protons: positively charged, located in the nucleus.
- Neutrons: neutral, located in the nucleus.
- Electrons: negatively charged, located in the electron cloud.

2. Label the following parts of the periodic table: s-block, p-block, d-block, f-block, transition metals, main group elements, noble gases, halogens, alkali metals, lanthanides, actinides.

3. Complete the following table:

Block	Groups	Periods	Elements
s-block	1, 2	1-7	Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra
p-block	13-18	2-7	B, C, N, O, F, Ne, Al, Si, P, S, Cl, Ar, As, Se, Br, Kr, Sb, Te, I, Xe, Bi, Po, At, Rn
d-block	3-10	4-7	Ti, V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Ga, Ge, As, Se, Br, Kr, Ag, Cd, In, Sn, Sb, Te, I, Xe, Pt, Au, Hg, Tl, Pb, Bi, Po, At, Rn
f-block	14-15	6-7	Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, Lu, Th, Pa, U, Np, Pu, Am, Cm, Bk, Cf, Es, Fm, Md, No, Lr

4. Explain the periodic trends of atomic radius, electronegativity, ionization energy, and electron affinity.

5. Complete the following table:

Element	Atomic Number	Electronegativity	Ionization Energy (kJ/mol)	Electron Affinity (kJ/mol)
Hydrogen	1	2.20	1312	-72
Helium	2	2.20	2372	0
Lithium	3	0.98	520	-60
Boron	5	2.04	801	-26
Carbon	6	2.55	1086	-122
Nitrogen	7	3.04	1402	-70
Oxygen	8	3.44	1314	-141
Fluorine	9	3.98	1681	-328
Neon	10	2.20	2081	0
Sodium	11	0.93	496	-48
Magnesium	12	1.31	738	-74
Aluminum	13	1.61	578	-42
Silicon	14	1.90	786	-112
Phosphorus	15	2.19	1012	-78
Sulfur	16	2.58	1000	-200
Chlorine	17	3.16	1251	-349
Argon	18	2.20	1521	0
Potassium	19	0.82	419	-48
Calcium	20	1.00	590	-29
Gallium	31	1.81	579	-29
Germanium	32	2.01	762	-119
Antimony	51	2.20	801	-103
Tellurium	52	2.36	864	-119
Iodine	53	2.50	1012	-101
Xenon	54	2.60	1171	-77
Bismuth	83	2.02	801	-103
Polonium	84	2.20	864	-119
Astatine	85	2.36	864	-119
Francium	87	0.70	402	-48
Radium	88	0.90	509	-29

6. Complete the following table:

Element	Atomic Number	Electronegativity	Ionization Energy (kJ/mol)	Electron Affinity (kJ/mol)
Hydrogen	1	2.20	1312	-72
Helium	2	2.20	2372	0
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Chlorine	17	3.16	1251	-349
Argon	18	2.20	1521	0
Potassium	19	0.82	419	-48
Calcium	20	1.00	590	-29
Gallium	31	1.81	579	-29
Germanium	32	2.01	762	-119
Antimony	51	2.20	801	-103
Tellurium	52	2.36	864	-119
Iodine	53	2.50	1012	-101
Xenon	54	2.60	1171	-77
Bismuth	83	2.02	80	

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**John Krenos, Joseph Potenza, Carl
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