

Gases and Moles

Date: _____

Hour: _____

Information: The Gas Constant

We previously examined the fact that $\frac{PV}{T}$ is a constant as long as the number of moles of a gas remains unchanged.

This means that any time you measure a gas, the pressure times the volume divided by the temperature (in Kelvin) will always equal the same number unless the amount of gas in the container changes. We are now going to explore how the pressure, volume and temperature depend on the number of moles of gas that are present.

Table 1: Experimental Data for Three Different Samples of Gas That Contain the Same Amount of Gas (# of moles of gas is constant)

Gas	# of moles of Gas in container (mol)	Volume of Container (L)	Temperature of Gas (K)	Pressure of Gas (kPa)
A	0.3008	14.0	672.0	120.0
B	0.3008	16.0	1280	200.0
C	0.3008	10.0	340.0	85.0

Table 2: Experimental Data for Three Different Samples of Gases That Contain Different Amounts of Gas (# of moles of gas is not constant)

Gas	# of moles of Gas in container	Volume of Container (L)	Temperature of Gas (K)	Pressure of Gas (kPa)
D	3.3485	40.0	230.0	160.0
E	0.7306	22.5	315	85
F	0.85950	16.0	280	125

Critical Thinking Questions

1. Prove that $\frac{PV}{T}$ is a constant for the data in Table 1, but not a constant for the data in Table 2.

2. Why is $\frac{PV}{T}$ a constant in Table 1, but not in Table 2?

Chemquest 36 Gases And Moles Answer Key

M Lipman



Chemquest 36 Gases And Moles Answer Key:

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- Chapter 2: Essential Elements of Chemquest 36 Gases And Moles Answer Key
- Chapter 3: Chemquest 36 Gases And Moles Answer Key in Everyday Life
- Chapter 4: Chemquest 36 Gases And Moles Answer Key in Specific Contexts
- Chapter 5: Conclusion

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Table of Contents Chemquest 36 Gases And Moles Answer Key

1. Understanding the eBook Chemquest 36 Gases And Moles Answer Key
 - The Rise of Digital Reading Chemquest 36 Gases And Moles Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemquest 36 Gases And Moles Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemquest 36 Gases And Moles Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemquest 36 Gases And Moles Answer Key
 - Personalized Recommendations
 - Chemquest 36 Gases And Moles Answer Key User Reviews and Ratings
 - Chemquest 36 Gases And Moles Answer Key and Bestseller Lists
5. Accessing Chemquest 36 Gases And Moles Answer Key Free and Paid eBooks
 - Chemquest 36 Gases And Moles Answer Key Public Domain eBooks
 - Chemquest 36 Gases And Moles Answer Key eBook Subscription Services
 - Chemquest 36 Gases And Moles Answer Key Budget-Friendly Options
6. Navigating Chemquest 36 Gases And Moles Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemquest 36 Gases And Moles Answer Key Compatibility with Devices
 - Chemquest 36 Gases And Moles Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemquest 36 Gases And Moles Answer Key
 - Highlighting and Note-Taking Chemquest 36 Gases And Moles Answer Key
 - Interactive Elements Chemquest 36 Gases And Moles Answer Key

8. Staying Engaged with Chemquest 36 Gases And Moles Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemquest 36 Gases And Moles Answer Key
9. Balancing eBooks and Physical Books Chemquest 36 Gases And Moles Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemquest 36 Gases And Moles Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemquest 36 Gases And Moles Answer Key
 - Setting Reading Goals Chemquest 36 Gases And Moles Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemquest 36 Gases And Moles Answer Key
 - Fact-Checking eBook Content of Chemquest 36 Gases And Moles Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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