

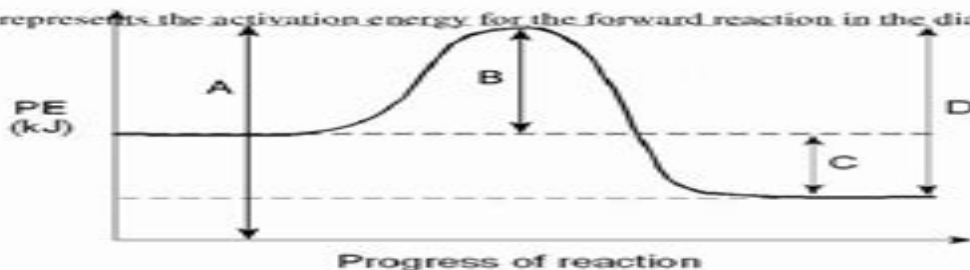
PART I
Total Value: 50%

Instructions: Shade the letter of the correct answer on the computer scorable answer sheet provided.

1. Why does increasing temperature increase reaction rate?

- (A) decreases activation energy
(B) decreases the percentage of effective collisions
✓ (C) increases activation energy
(D) increases the percentage of effective collisions

2. Which represents the activation energy for the forward reaction in the diagram below?



- ✓ (A) A
(B) B
(C) C
(D) D

3. Which type of change is best for monitoring the reaction rate of the reaction below?



- (A) colour
✓ (B) mass
(C) pH
(D) volume

4. Which describes an activated complex?

- (A) has a lower potential energy than products
(B) has a lower potential energy than reactants
✓ (C) stable species that forms upon collision of reactants
(D) unstable species that forms upon collision of reactants

5. According to the collision theory, why does the reaction below occur in more than one step?



- ✓ (A) low $\text{C}_5\text{H}_{12}(\text{g})$ concentration
(B) low probability of a nine-particle collision
(C) particles collide with insufficient kinetic energy
(D) temperature is very high

Collision Theory Answer Key

Robert R. Redfield

A red, semi-circular logo with a gradient, resembling a stylized 'R' or a drop, positioned to the right of the author's name.

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Table of Contents Collision Theory Answer Key

1. Understanding the eBook Collision Theory Answer Key
 - The Rise of Digital Reading Collision Theory Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Collision Theory Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Collision Theory Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Collision Theory Answer Key
 - Personalized Recommendations

- Collision Theory Answer Key User Reviews and Ratings
- Collision Theory Answer Key and Bestseller Lists
- 5. Accessing Collision Theory Answer Key Free and Paid eBooks
 - Collision Theory Answer Key Public Domain eBooks
 - Collision Theory Answer Key eBook Subscription Services
 - Collision Theory Answer Key Budget-Friendly Options
- 6. Navigating Collision Theory Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Collision Theory Answer Key Compatibility with Devices
 - Collision Theory Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Collision Theory Answer Key
 - Highlighting and Note-Taking Collision Theory Answer Key
 - Interactive Elements Collision Theory Answer Key
- 8. Staying Engaged with Collision Theory Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Collision Theory Answer Key
- 9. Balancing eBooks and Physical Books Collision Theory Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Collision Theory Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Collision Theory Answer Key
 - Setting Reading Goals Collision Theory Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Collision Theory Answer Key
 - Fact-Checking eBook Content of Collision Theory Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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