

ANSWERS:

1. **What are acids and bases, and how are they characterized?**

- Acids are substances that can donate a proton (H^+ ion) to another substance, characterized by their sour taste, ability to turn blue litmus paper red, and corrosive nature. Bases are substances that can accept a proton (H^+ ion) from another substance, known for their bitter taste, ability to turn red litmus paper blue, and slippery feel.

2. **What are some common examples of acids and bases?**

- Common examples of acids include hydrochloric acid (HCl) and citric acid, while examples of bases include sodium hydroxide (NaOH) and ammonia (NH_3).

3. **What is the pH scale, and how does it measure acidity or basicity?**

- The pH scale measures the acidity or basicity of a solution on a scale from 0 to 14. A pH of 7 is considered neutral, while solutions with a pH below 7 are acidic, and solutions with a pH above 7 are basic.

4. **Describe the process of neutralization and provide an example.**

- Neutralization is the process in which acids and bases react with each other to form water and a salt. An example is the reaction between hydrochloric acid and sodium hydroxide, producing water and sodium chloride: $\text{HCl} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaCl}$.

5. **In what fields is understanding acids and bases important?**

- Understanding acids and bases is important in various fields, including medicine, agriculture, and manufacturing.

6. **Why is knowledge of acid-base balance crucial in medicine?**

- Knowledge of acid-base balance is crucial in medicine for diagnosing and treating diseases because it helps maintain the body's internal environment within a narrow pH range, which is essential for proper cellular function and enzyme activity.

Chapter 20 Acids Bases Guided Reading Answers

LL Leslie

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