

1. In the list below, circle the properties of acids and box the properties of bases.

Sour taste	Feel slippery	Contains hydroxide ions
Turns litmus blue	Reacts with metals to form H_2	Turns litmus red
Bitter taste	Contains hydrogen ions	Corrosive

2. Complete the chart below:

Acid-Base Definitions		
Type	Acid Definition	Base Definition
Arrhenius	Donates an H^+ ion in solution	Donates an OH^- ion in solution
Bronsted-Lowry	Donates an H^+ ion	Accepts an H^+ ion
Lewis	Accepts an electron pair	Donates an electron pair

3. What term(s) describe a substance that can act as an acid or a base? **Amphoteric or amphiprotic**
4. Neutralization reactions always produce a **SALT** and **WATER**
5. If the $[H^+]$ is greater than the $[OH^-]$, the solution is acidic basic neutral
6. If the $[OH^-]$ is greater than the $[H^+]$, the solution is acidic basic neutral
7. If the $[H^+]$ is equal to the $[OH^-]$, the solution is acidic basic neutral
8. If the pH is 7, the solution is acidic basic neutral
9. If the pH is greater than 7, the solution is acidic basic neutral
10. If the pH is less than 7, the solution is acidic basic neutral
11. How many moles of lithium hydroxide are required to neutralize 19.2 grams of chloric acid?



$$(19.2 \text{ g } HClO_3) \times (1 \text{ mol } HClO_3/84 \text{ g } HClO_3) \times (1 \text{ mol } LiOH/1 \text{ mol } HClO_3) = \underline{0.228 \text{ mol } LiOH}$$

12. How many grams of calcium hydroxide are required to neutralize 132.4 grams of nitric acid?



$$(132.4 \text{ g } HNO_3) \times (1 \text{ mol } HNO_3/63 \text{ g } HNO_3) \times (1 \text{ mol } Ca(OH)_2/2 \text{ mol } HNO_3) \times (74 \text{ g } Ca(OH)_2/1 \text{ mol } Ca(OH)_2) = \underline{77.76 \text{ g } Ca(OH)_2}$$

Chapter 19 Acids And Bases Answers

David R. Klein



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Table of Contents Chapter 19 Acids And Bases Answers

1. Understanding the eBook Chapter 19 Acids And Bases Answers
 - The Rise of Digital Reading Chapter 19 Acids And Bases Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 19 Acids And Bases Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 19 Acids And Bases Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 19 Acids And Bases Answers
 - Personalized Recommendations

- Chapter 19 Acids And Bases Answers User Reviews and Ratings
- Chapter 19 Acids And Bases Answers and Bestseller Lists
- 5. Accessing Chapter 19 Acids And Bases Answers Free and Paid eBooks
 - Chapter 19 Acids And Bases Answers Public Domain eBooks
 - Chapter 19 Acids And Bases Answers eBook Subscription Services
 - Chapter 19 Acids And Bases Answers Budget-Friendly Options
- 6. Navigating Chapter 19 Acids And Bases Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 19 Acids And Bases Answers Compatibility with Devices
 - Chapter 19 Acids And Bases Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 19 Acids And Bases Answers
 - Highlighting and Note-Taking Chapter 19 Acids And Bases Answers
 - Interactive Elements Chapter 19 Acids And Bases Answers
- 8. Staying Engaged with Chapter 19 Acids And Bases Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 19 Acids And Bases Answers
- 9. Balancing eBooks and Physical Books Chapter 19 Acids And Bases Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 19 Acids And Bases Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 19 Acids And Bases Answers
 - Setting Reading Goals Chapter 19 Acids And Bases Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 19 Acids And Bases Answers
 - Fact-Checking eBook Content of Chapter 19 Acids And Bases Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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