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Post-Laboratory Questions

(Use the spaces provided for the answers and additional paper if necessary.)

- A student doing the experiment in this module extended the study of the [CoCl₆]²⁻ ion/[Co(l₂Ol₆]²⁻ ion equilibrium. In one test tube, the student added silver nitrate (AgNO₂) to a blue equilibrium mixture prepared from CoCl₂. The reaction mixture became pink and cloudy, and a white precipitate settled out leaving a clear, pink solution. The student identified the precipitate as silver chloride, AgCl.
 - (1) Briefly explain how these observations are consistent with Le Châtelier's principle.

In a second test, the student placed a test tube containing a pink equilibrium mixture in a hot-water bath. The solution turned blue. When the student removed the test tube from the hot-water bath and placed it in an ice-water bath, the solution turned pink.

(2) Is the forward reaction in the [CoCl₄]²- ion/[Co(H₂O)₆]²* ion equilibrium exothermic or endothermic?

(3) Briefly explain how the student's observations support your answer to (2).

(4) Write the net ionic equation for this equilibrium, including heat.

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Robert R. Redfield

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