

Name: _____

Chemistry Unit 3

Redox (No Electrolysis) Progress Test

- redox reactions with reference to electron transfer, reduction and oxidation reactions, reducing and oxidising agents, and use of oxidation numbers to identify conjugate reducing and oxidising agents
- the writing of balanced half-equations for oxidation and reduction reactions and balanced ionic equations, including states, for overall redox reactions
- galvanic cells as primary cells and as portable or fixed chemical energy storage devices that can produce electricity (details of specific cells not required) including common design features (anode, cathode, electrolytes, salt bridge and separation of half-cells) and chemical processes (electron and ion flows, half-equations and overall equations)
- the comparison of the energy transformations occurring in spontaneous exothermic redox reactions involving direct contact between reactants (transformation of chemical energy to heat energy) compared with those occurring when the reactants are separated in galvanic cells (transformation of chemical energy to electrical energy)
- the use of the electrochemical series in designing and constructing galvanic cells and as a tool for predicting the products of redox reactions, deducing overall equations from redox half-equations and determining maximum cell voltage under standard conditions.
- the common design features of fuel cells including use of porous electrodes for gaseous reactants to increase cell efficiency (details of specific cells not required)
- the comparison of the use of fuel cells and combustion of fuels to supply energy with reference to their energy efficiencies (qualitative), safety, fuel supply (including the storage of hydrogen), production of greenhouse gases and applications
- the comparison of fuel cells and galvanic cells with reference to their definitions, functions, design features, energy transformations, energy efficiencies (qualitative) and applications.

SECTION A - Multiple-choice questions (10 marks)

Please mark your responses (A, B, C or D) to each of the multiple-choice questions in the table below.

1.	2.	3.	4.	5.
6.	7.	8.	9.	10.

Question 1

Magnesium reacts with water to slowly form hydrogen and magnesium hydroxide. The reaction may be represented by the equation below.



During this reaction, water is acting as

- A. An acid
- B. A base
- C. A reducing agent
- D. An oxidising agent

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