

Chemistry Study Guide

Periodic Table of the Elements

¹ ₁ H																	² ₂ He																														
³ ₃ Li	⁴ ₄ Be																	⁵ ₅ B	⁶ ₆ C	⁷ ₇ N	⁸ ₈ O	⁹ ₉ F	¹⁰ ₁₀ Ne																								
¹¹ ₁₁ Na	¹² ₁₂ Mg																	¹³ ₁₃ Al	¹⁴ ₁₄ Si	¹⁵ ₁₅ P	¹⁶ ₁₆ S	¹⁷ ₁₇ Cl	¹⁸ ₁₈ Ar																								
¹⁹ ₁₉ K	²⁰ ₂₀ Ca	²¹ ₂₁ Sc	²² ₂₂ Ti	²³ ₂₃ V	²⁴ ₂₄ Cr	²⁵ ₂₅ Mn	²⁶ ₂₆ Fe	²⁷ ₂₇ Co	²⁸ ₂₈ Ni	²⁹ ₂₉ Cu	³⁰ ₃₀ Zn	³¹ ₃₁ Ga	³² ₃₂ Ge	³³ ₃₃ As	³⁴ ₃₄ Se	³⁵ ₃₅ Br	³⁶ ₃₆ Kr																														
³⁷ ₃₇ Rb	³⁸ ₃₈ Sr	³⁹ ₃₉ Y	⁴⁰ ₄₀ Zr	⁴¹ ₄₁ Nb	⁴² ₄₂ Mo	⁴³ ₄₃ Tc	⁴⁴ ₄₄ Ru	⁴⁵ ₄₅ Rh	⁴⁶ ₄₆ Pd	⁴⁷ ₄₇ Ag	⁴⁸ ₄₈ Cd	⁴⁹ ₄₉ In	⁵⁰ ₅₀ Sn	⁵¹ ₅₁ Sb	⁵² ₅₂ Te	⁵³ ₅₃ I	⁵⁴ ₅₄ Xe																														
⁵⁵ ₅₅ Cs	⁵⁶ ₅₆ Ba			⁷² ₇₂ Hf	⁷³ ₇₃ Ta	⁷⁴ ₇₄ W	⁷⁵ ₇₅ Re	⁷⁶ ₇₆ Os	⁷⁷ ₇₇ Ir	⁷⁸ ₇₈ Pt	⁷⁹ ₇₉ Au	⁸⁰ ₈₀ Hg	⁸¹ ₈₁ Tl	⁸² ₈₂ Pb	⁸³ ₈₃ Bi	⁸⁴ ₈₄ Po	⁸⁵ ₈₅ At	⁸⁶ ₈₆ Rn																													
⁸⁷ ₈₇ Fr	⁸⁸ ₈₈ Ra			¹⁰² ₁₀₂ Rf	¹⁰³ ₁₀₃ Db	¹⁰⁴ ₁₀₄ Sg	¹⁰⁵ ₁₀₅ Bh	¹⁰⁶ ₁₀₆ Hs	¹⁰⁷ ₁₀₇ Mt	¹⁰⁸ ₁₀₈ Ds	¹⁰⁹ ₁₀₉ Rg	¹¹⁰ ₁₁₀ Cn	¹¹¹ ₁₁₁ Uut	¹¹² ₁₁₂ Fl	¹¹³ ₁₁₃ Uup	¹¹⁴ ₁₁₄ Lv	¹¹⁵ ₁₁₅ Uus	¹¹⁶ ₁₁₆ Uuo																													
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⁵⁷ ₅₇ La	⁵⁸ ₅₈ Ce	⁵⁹ ₅₉ Pr	⁶⁰ ₆₀ Nd	⁶¹ ₆₁ Pm	⁶² ₆₂ Sm	⁶³ ₆₃ Eu	⁶⁴ ₆₄ Gd	⁶⁵ ₆₅ Tb	⁶⁶ ₆₆ Dy	⁶⁷ ₆₇ Ho	⁶⁸ ₆₈ Er	⁶⁹ ₆₉ Tm	⁷⁰ ₇₀ Yb	⁷¹ ₇₁ Lu																																	
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Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

Chemistry Study Guide Review Answers

Peter L. Kresan, Reed Mencke



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Table of Contents Chemistry Study Guide Review Answers

1. Understanding the eBook Chemistry Study Guide Review Answers
 - The Rise of Digital Reading Chemistry Study Guide Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide Review Answers
 - Personalized Recommendations
 - Chemistry Study Guide Review Answers User Reviews and Ratings
 - Chemistry Study Guide Review Answers and Bestseller Lists

5. Accessing Chemistry Study Guide Review Answers Free and Paid eBooks
 - Chemistry Study Guide Review Answers Public Domain eBooks
 - Chemistry Study Guide Review Answers eBook Subscription Services
 - Chemistry Study Guide Review Answers Budget-Friendly Options
6. Navigating Chemistry Study Guide Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Study Guide Review Answers Compatibility with Devices
 - Chemistry Study Guide Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Study Guide Review Answers
 - Highlighting and Note-Taking Chemistry Study Guide Review Answers
 - Interactive Elements Chemistry Study Guide Review Answers
8. Staying Engaged with Chemistry Study Guide Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide Review Answers
9. Balancing eBooks and Physical Books Chemistry Study Guide Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Study Guide Review Answers
 - Setting Reading Goals Chemistry Study Guide Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Study Guide Review Answers
 - Fact-Checking eBook Content of Chemistry Study Guide Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
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