

Limiting Reagent & Percent Yield Practice Worksheet

1. When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

a. Write the balanced equation for the reaction given above:



- b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.362 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol CuCl}_2} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13 \text{ g NaCl}$$

$$20 \text{ g NaNO}_3 \times \frac{1 \text{ mol NaNO}_3}{84.997 \text{ g NaNO}_3} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol NaNO}_3} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13.8 \text{ g NaCl}$$

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? **copper (II) chloride**

- d. How many grams of copper (II) nitrate is formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{1 \text{ mol Cu(NO}_3)_2}{1 \text{ mol CuCl}_2} \times \frac{187.467 \text{ g Cu(NO}_3)_2}{1 \text{ mol Cu(NO}_3)_2} = 20.9 \text{ g Cu(NO}_3)_2$$

- e. How much of the excess reagent is left over in this reaction?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol CuCl}_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 19 \text{ g NaNO}_3$$

$$20 \text{ g NaNO}_3 - 19 \text{ g NaNO}_3 = 1 \text{ g NaNO}_3$$

- f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$$

$$11.3/13.0 \times 100\% = 86.9\%$$

2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.

a. Balance the following equation:



- b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

$$25.0 \text{ g Pb(NO}_3)_2 \times \frac{1 \text{ mol Pb(NO}_3)_2}{331.214 \text{ g Pb(NO}_3)_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol Pb(NO}_3)_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 12.8 \text{ g NaNO}_3$$

$$15.0 \text{ g NaI} \times \frac{1 \text{ mol NaI}}{149.89 \text{ g NaI}} \times \frac{2 \text{ mol NaNO}_3}{2 \text{ mol NaI}} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

8.51 g NaNO₃ can be made

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1. Understanding the eBook Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key
 - The Rise of Digital Reading Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key
 - User-Friendly Interface

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4. Exploring eBook Recommendations from Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key
 - Personalized Recommendations
 - Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key User Reviews and Ratings
 - Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key and Bestseller Lists
5. Accessing Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key Free and Paid eBooks
 - Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key Public Domain eBooks
 - Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key eBook Subscription Services
 - Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key Budget-Friendly Options
6. Navigating Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key Compatibility with Devices
 - Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key
 - Highlighting and Note-Taking Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key
 - Interactive Elements Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key
8. Staying Engaged with Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key
9. Balancing eBooks and Physical Books Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key
 - Setting Reading Goals Chemistry Worksheet 12 3 Limiting Reagent And Pencent Yield With Anser Key

- Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Worksheet 12 3 Limiting Reagent And Percent Yield With Answer Key
 - Fact-Checking eBook Content of Chemistry Worksheet 12 3 Limiting Reagent And Percent Yield With Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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