

1. Here are three isotopes of an element: $^{12}_6\text{C}$ $^{13}_6\text{C}$ $^{14}_6\text{C}$
 - a. The element is: **Carbon**
 - b. The number 6 refers to the: **Atomic Number**
 - c. The numbers 12, 13, and 14 refer to the **different mass numbers**
 - d. How many protons and neutrons are in the first isotope? **P: 6 N: 6**
 - e. How many protons and neutrons are in the second isotope? **P: 6 N: 7**
 - f. How many protons and neutrons are in the third isotope? **P: 6 N: 8**

2. Complete the following chart:

Isotope name	atomic #	mass #	# of protons	# of neutrons	# of electrons
Potassium-37	19	37	19	18	19
Oxygen-17	8	17	8	9	8
uranium-235	92	235	92	143	92
uranium-238	92	238	92	146	92
boron-10	5	10	5	5	5
boron-11	5	11	5	6	5

3. How do you calculate the mass number of an element?

Mass # = protons + neutrons

4. If an atom has 14 protons, 17 neutrons and 14 electrons, which expression identifies the element?
 - a. Silicon-14
 - b. Silicon-28
 - c. **Silicon-31**
 - d. Silicon-17
5. If an atom has 53 protons, 54 neutrons and 53 electrons, which expression identifies the element?
 - a. Iodine-53
 - b. Iodine-127
 - c. Iodine-54
 - d. **Iodine-107**

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Rachel Turoscy



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Table of Contents Compass Learning Odyssey Answer Key Chemistry

1. Understanding the eBook Compass Learning Odyssey Answer Key Chemistry
 - The Rise of Digital Reading Compass Learning Odyssey Answer Key Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Compass Learning Odyssey Answer Key Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Compass Learning Odyssey Answer Key Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Compass Learning Odyssey Answer Key Chemistry
 - Personalized Recommendations
 - Compass Learning Odyssey Answer Key Chemistry User Reviews and Ratings
 - Compass Learning Odyssey Answer Key Chemistry and Bestseller Lists
5. Accessing Compass Learning Odyssey Answer Key Chemistry Free and Paid eBooks

- Compass Learning Odyssey Answer Key Chemistry Public Domain eBooks
 - Compass Learning Odyssey Answer Key Chemistry eBook Subscription Services
 - Compass Learning Odyssey Answer Key Chemistry Budget-Friendly Options
6. Navigating Compass Learning Odyssey Answer Key Chemistry eBook Formats
- ePub, PDF, MOBI, and More
 - Compass Learning Odyssey Answer Key Chemistry Compatibility with Devices
 - Compass Learning Odyssey Answer Key Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Compass Learning Odyssey Answer Key Chemistry
 - Highlighting and Note-Taking Compass Learning Odyssey Answer Key Chemistry
 - Interactive Elements Compass Learning Odyssey Answer Key Chemistry
8. Staying Engaged with Compass Learning Odyssey Answer Key Chemistry
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Compass Learning Odyssey Answer Key Chemistry
9. Balancing eBooks and Physical Books Compass Learning Odyssey Answer Key Chemistry
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Compass Learning Odyssey Answer Key Chemistry
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Compass Learning Odyssey Answer Key Chemistry
- Setting Reading Goals Compass Learning Odyssey Answer Key Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Compass Learning Odyssey Answer Key Chemistry
- Fact-Checking eBook Content of Compass Learning Odyssey Answer Key Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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