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**Charles Henry Edwards, David E.
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web write the oxidation and reduction reactions for each redox reaction the first one is done for you $23 \text{ Fe} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$
oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$ reduction $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3 \text{ Ag}$ oxidation reduction $25 \text{ Cu} + 2 \text{ Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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web in each of the following redox reactions indicate the elements that undergo changes in oxidation number and the magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a $\text{I}_2 + \text{O}_5$

5co i2 5co2 i 2 o 5 5 c o i 2 5 c o 2

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web chemical reactions in which there is a transfer of electrons from one substance to another are known as oxidation reduction reactions or redox reactions in this unit you will examine the oxidation reduction process and use the oxidation state and oxidation number concepts to identify redox reactions and to keep track of electrons

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also $2\text{sr} + \text{o}_2 \rightarrow 2\text{sro}$ sr^0 to sr^2 oxidized reducing agent o^0 to o^2 reduced ox ag $2\text{li} + \text{s} \rightarrow \text{li}_2\text{s}$ li^0 to li^1 oxidized red ag s^0 to s^2 reduced ox ag $3\text{cs} + 2\text{br}_2 \rightarrow 3\text{csbr}_2$