NATURAL SCIENCES

Gr. 9 - LESSON PLAN - MATTER & MATERIALS

	TOPIC: COMPOU	INDS	
Sub-topic: Names of Compounds		Duration: 1 Hour	
	CONTENT & CONCEPTS	(CAPS p63)	
KEY CONCEPTS			
1. Chemical	reaction		
Chemical	formula		
Names of	compounds according to their elements.		
Specific Aims:	Specific Aim 1: 'Doing Science'		X
Specific Airis.			
	Specific Aim 2: 'Knowing the subject content and making connections'		X
	Specific Aim 3: 'Understanding the uses	of Science'	

LESSON OBJECTIVES:

- Learners will be able to write the names of compounds as indicated on page 63 of the CAPS Document.
- 2. Learners will be able to make models of different compounds

RESOURCES REQUIRED:

- Periodic Table Poster
- Siyavula Sasol Inzalo (Learner book) page 163

TEACHING & LEARNING ACTIVITY: Introduction Briefly review and revise concepts dealt with in Grade 8, focusing on compounds

As each element has a unique name for example:

- Hydrogen
- Lithium

Remember the first 20 elements learnt in grade 8. How did you memorize them? Do you remember Mnemonics...

Hit - H; He - He; Lies - Li; Because - Be; Boys - B; Can - C; Not - N; Operate - O; Fireplaces - F;

New - No: Nation - No: Might - Mg; Also - Al; Sign - Si; Peace - P; Security - S; Clause - Cl

A - Ar; King - K; Can - Ca LEARNERS MAY RECITE THIS MINEMONIC AS A WAY OF MEMORIZING THE FIRST 20 ELEMENTS.

The teacher provides the following explanation to the learners

Each compound has a unique name. How is a compound formed?

- A compound is formed when TWO or more different elements combine chemically.
- Each atom in a compound are held by chemical bonds
- Compounds are formed during chemical reaction (when elements/ componds combine to form new substances)
- Each compound can be written as a **formula**, for example.

 H-O: NaCl: CuSO₄

What is a formula?

- This is a notation that shows which elements the compound is made up of.
- It indicates the total number of atoms in one unit of the compound
- It indicates the ratio in which atoms of different elements are bonded to each other.
 It gives the total number of atoms in one unit of a compound.

For example, the formula for Copper Sulphate is CuSO.

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Lei Shi

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Table of Contents Caps Lesson Plan Grade 9 Natural Sciences

- 1. Understanding the eBook Caps Lesson Plan Grade 9 Natural Sciences
 - The Rise of Digital Reading Caps Lesson Plan Grade 9 Natural Sciences
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Caps Lesson Plan Grade 9 Natural Sciences
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Caps Lesson Plan Grade 9 Natural Sciences
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Caps Lesson Plan Grade 9 Natural Sciences
 - Personalized Recommendations
 - Caps Lesson Plan Grade 9 Natural Sciences User Reviews and Ratings
 - Caps Lesson Plan Grade 9 Natural Sciences and Bestseller Lists
- 5. Accessing Caps Lesson Plan Grade 9 Natural Sciences Free and Paid eBooks
 - Caps Lesson Plan Grade 9 Natural Sciences Public Domain eBooks
 - Caps Lesson Plan Grade 9 Natural Sciences eBook Subscription Services
 - Caps Lesson Plan Grade 9 Natural Sciences Budget-Friendly Options
- 6. Navigating Caps Lesson Plan Grade 9 Natural Sciences eBook Formats

- o ePub, PDF, MOBI, and More
- Caps Lesson Plan Grade 9 Natural Sciences Compatibility with Devices
- Caps Lesson Plan Grade 9 Natural Sciences Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Caps Lesson Plan Grade 9 Natural Sciences
 - Highlighting and Note-Taking Caps Lesson Plan Grade 9 Natural Sciences
 - Interactive Elements Caps Lesson Plan Grade 9 Natural Sciences
- 8. Staying Engaged with Caps Lesson Plan Grade 9 Natural Sciences
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Caps Lesson Plan Grade 9 Natural Sciences
- 9. Balancing eBooks and Physical Books Caps Lesson Plan Grade 9 Natural Sciences
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Caps Lesson Plan Grade 9 Natural Sciences
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Caps Lesson Plan Grade 9 Natural Sciences
 - $\circ\,$ Setting Reading Goals Caps Lesson Plan Grade 9 Natural Sciences
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Caps Lesson Plan Grade 9 Natural Sciences
 - Fact-Checking eBook Content of Caps Lesson Plan Grade 9 Natural Sciences
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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