

Date Monday 10/18/2021 Name Aya Turner
Section _____ Team _____
Instructor _____

Pre-Lab Study Questions

1. What is meant by the simplest formula of a compound?

The lowest whole number ratio of the atoms in the formula.

2. How does a hydrate differ from an anhydrate?

Hydrate is an ionic compound with water and anhydrate is an ionic compound without water.

3. What happens when a hydrate is heated?

When a hydrate is heated, it has the water removed and becomes an anhydrate.

4. A hydrate of CoCl_2 with a mass of 6.00 g is heated strongly. After cooling, the mass of the anhydrate is 3.27 g.

- a. How many grams of H_2O were lost from the hydrate?

$$\begin{array}{r} 6.00 \\ - 3.27 \\ \hline 2.73 \end{array} \quad 2.73 \text{ g } \text{H}_2\text{O}$$

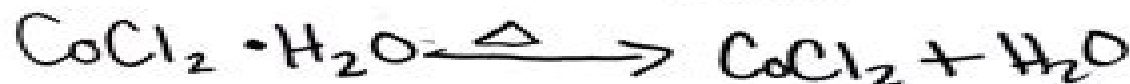
- b. What is the % water in the hydrate?

$$\frac{2.73}{6.00} \times 100 = 45.5\% \text{ of } \text{H}_2\text{O}$$

- c. What is the formula of the CoCl_2 hydrate?



- d. Write the equation for the dehydration of the CoCl_2 hydrate.



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Denise Guinn, Rebecca Brewer



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