

Stoichiometry Study Guide KEY

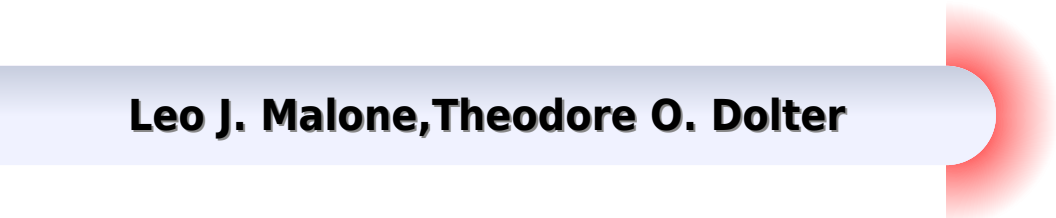
Chemistry

RHS – Mr. Moss

1. Define the following:
 - a. *Stoichiometry-the study of the quantitative relationships between the amounts of reactants used and the products formed by a chemical reaction.*
 - b. *Mole –The SI unit used to measure the amount of a substance that contains 6.02×10^{23} atoms of that substance.*
 - c. *Mole Ratio-The ratio between any two substances in a balanced chemical equation.*
 - d. *Molar Mass-The mass in grams of one mole of any substance.*
2. What relationships can be determined from a balanced chemical equation? *The relationships between particles, moles, and mass for all reactants and products.*
3. Explain why mole ratios are central to stoichiometric calculations. *Mole ratios allow for the conversion from moles of one substance in a balanced chemical equation to moles of another substance in the same equation.*
4. What is the first step in all stoichiometric calculation? (Be careful.) *Write a balanced chemical equation.*
5. How is molar mass used in some stoichiometric calculations? *Molar mass is a conversion factor for converting moles of a given substance to mass or mass of a given substance to moles.*
6. Find the molar mass of the following:
 - a. H_2SO_4 (sulfuric acid) *98.09 g/mol*
 - b. ZnO (zinc oxide) *81.39 g/mol*
 - c. NH_4NO_3 (ammonium nitrate) *80.05 g/mol*
7. List the mole ratios of the following *balanced* chemical equation.
 - a. $2\text{Al} + 3\text{Br}_2 \rightarrow 2\text{AlBr}_3$ $\frac{2 \text{ mol Al}}{3 \text{ mol Br}_2} ; \frac{3 \text{ mol Br}_2}{2 \text{ mol AlBr}_3} ; \frac{2 \text{ mol Al}}{2 \text{ mol AlBr}_3}$ *(Each one of these can be inverted to form another mole ratio. This would give a total of 6 ratios.)*

Chemistry Stoichiometry Study Guide Answer Key

Leo J. Malone, Theodore O. Dolter



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Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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