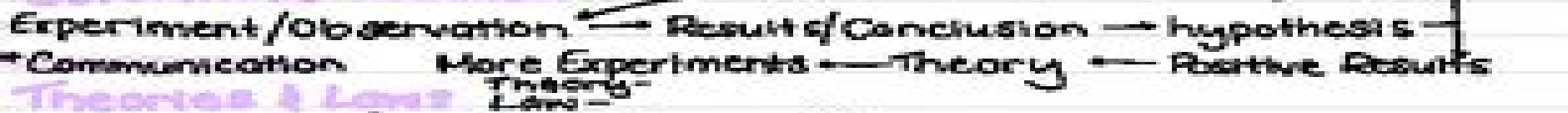


Finals Grinding

Chapter 1

1.1 Chemistry in Context

Scientific Method



1.2 Phases & Matter Classification

States of Matter

Solid - fixed shape and volume
 Liquid - shape of container, fixed volume
 gas - expands to fill container
 plasma - gaseous state w/ charged particles w/ unique properties

Mass & Weight

Mass - amt of matter
 Weight - force of gravity on object
 Conservation - Amt of matter doesn't change due to chemical or physical reactions

Defined Terms

Pure substance - constant composition
 Mixture - two or more substances
 Elements - same type of atom, smallest version
 Compounds - can be decomposed to other compounds/elements
 Homogeneous mixture - uniform composition
 Heterogeneous mixture - physically distinct parts

Miscible - forms a homogenous mixture when mixed

1.3 Physical & Chemical Properties

Ex/Intensive Properties

extensive depend on amt
 intensive independent of amt

Periodic Table

Atomic number = protons
 Periods - across
 Groups - down
 Chalcogens - group 16 (O, S, Se, Te, Po, Lv)
 Halogens - group 17 (F, Cl, Br, I, At, Ts)
 Alkali - group 1 (Li, Na, K, Rb, Cs, Fr)
 Alkaline earth - group 2 (Be, Mg, Ca, Sr, Ba, Ra)
 Transition - groups 3-10
 Metals - left side
 Metalloids - diagonal line
 Non - right side
 Halogens - group 17
 Noble gases - group 18

1.4 Measurements

magnitude - size of number
 Unit - standard of comparison
 Indication of uncertainty - $\pm$ one uncertain digit
 Pico 10^{-12} nano 10^{-9} micro 10^{-6} milli 10^{-3} centi 10^{-2} deci 10^{-1}
 Kilo 10^3 mega 10^6
 Density - mass/volume (intrinsic property)

1.5 Uncertainty, Accuracy, Precision

measurement device & user introduce uncertainty
 measure from bottom of meniscus

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