

Burning gunpowder transfers energy to the atoms of color-producing chemicals, exciting their electrons to higher energy states. In returning to the ground state, the electrons emit light of specific colors, resulting in spectacular fireworks displays. Strontium produces red, and sodium produces yellow/orange.

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28

Atomic Physics

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A hot gas emits light of certain characteristic wavelengths that can be used to identify it, much as a fingerprint can identify a person. For a given atom, these characteristic emitted wavelengths can be understood using physical quantities called quantum numbers. The simplest atom is hydrogen, and understanding it can lead to understanding the structure of other atoms and their combinations. The fact that no two electrons in an atom can have the same set of quantum numbers—the Pauli exclusion principle—is extremely important in understanding the properties of complex atoms and the arrangement of elements in the periodic table. Knowledge of atomic structure can be used to describe the mechanisms involved in the production of x-rays and the operation of a laser, among many other applications.

28.1 Early Models of the Atom

The model of the atom in the days of Newton was a tiny, hard, indestructible sphere. Although this model was a good basis for the kinetic theory of gases, new models had to be devised when later experiments revealed the electronic nature of atoms. J. J. Thomson (1856–1940) suggested a model of the atom as a volume of positive charge with electrons embedded throughout the volume, much like the seeds in a watermelon (Fig. 28.1).

In 1911 Ernest Rutherford (1871–1937) and his students Hans Geiger and Ernest Marsden performed a critical experiment showing that Thomson's model couldn't be correct. In this experiment a beam of positively charged **alpha particles** was projected against a thin metal foil, as in Figure 28.2a. Most of the alpha particles passed through the foil as if it were empty space, but a few particles were scattered through large angles, some even traveling backward.

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