

PROBLEMS WITH CALCULATIONS

1. Nitrogen gas and hydrogen gas combine according to the balanced equation below:



The diagram below shows an initial mixture of nitrogen gas (blue molecules) and hydrogen gas (red spheres) which react to form ammonia gas (blue spheres). Before the reaction, the total number of molecules is equal to the number of molecules after the reaction.



Questions:

1. How many molecules of N_2 are there? How many molecules of H_2 are there? How many molecules of NH_3 are there? How many molecules of N_2 are there? How many molecules of H_2 are there? How many molecules of NH_3 are there? How many molecules of N_2 are there? How many molecules of H_2 are there? How many molecules of NH_3 are there?

2. Use the same equation as question 1 to answer the following questions:

- a. If 10.0 g of N_2 and 1.0 g of H_2 are reacted, what mass of NH_3 is produced?

$$\begin{aligned} \text{N}_2 + 3\text{H}_2 &\rightarrow 2\text{NH}_3 \\ 28.0 \text{ g} &+ 6.0 \text{ g} \rightarrow 34.0 \text{ g} \\ 10.0 \text{ g} &+ 1.0 \text{ g} \rightarrow 11.0 \text{ g} \end{aligned}$$

- b. What is the limiting reagent? (showing work)

Answer: H_2 is the limiting reagent.

Answer: H_2

Answer: H_2

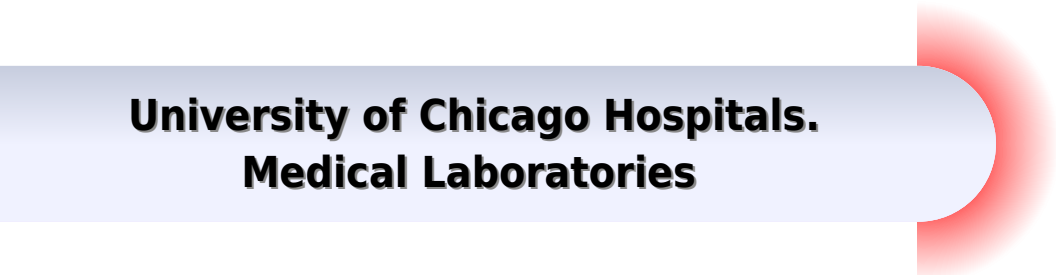
- c. What is the percent yield of NH_3 if 10.0 g of NH_3 is produced in this reaction?

$$\frac{10.0 \text{ g}}{11.0 \text{ g}} \times 100\%$$

Answer: 90.9%

Ccsf Chem 40 Lab Manual

**University of Chicago Hospitals.
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