

Concentration and Molarity PhET Labs

Name: _____

Procedure: https://phet.colorado.edu/sims/html/concentration/latest/concentration_en.html

Part 1: Dissolution and Saturation

Take some time to play and familiarize yourself with the simulation. Click on everything. Move all the sliders. Notice what happens to the concentration as solid solute is added and when evaporation occurs.



Concentration

Concentration
(mol/L)
0.179

- How does the concentration change as solid solute is added? concentration increases
 - How does the concentration change as additional water is added? concentration decreases
 - How does the concentration change as evaporation occurs? concentration increases
 - How do you know when a solution is saturated? concentration does not change (it also shows on screen)
 - When a solution is saturated, and additional solid solute is added, what happens? solid does not dissolve and settles at bottom
 - Why do you think this is? solvent can't hold any more solute
 - How does adding this additional solute change the concentration of this saturated solution? decreases concentration
 - How does evaporation change the concentration of a saturated solution? no change- solution is already saturated
- Part 2: Concentrated Solutions (same website as part 1 just click on the Solution option)**
10. Adding a concentrated solution ... describe a way to determine the concentration of the solution in the spigot. Write your plan here: empty the container and add only one solution from the dropper and then use the sensor to measure concentration

Solution



- Using your plan ... how might you get that concentrated solution to become saturated? you can evaporate
- Does your plan work for all the other solutions too? yes Why? / Why Not? evaporation reduces the amount of water (or solute) so concentration increases



Part 3: Molarity https://phet.colorado.edu/sims/html/molarity/latest/molarity_en.html

Molarity (M) = $\frac{\text{moles of solute (mol)}}{\text{volume of solution (L)}}$

Molarity is moles per liter, that is, how many moles of solute (entire salt) is dissolved per liter of solution.

First, determine the saturation concentration of each of the solutions, that is, how concentrated can you get each solution before the solution is saturated. If you can't determine the concentration using the simulation "Molarity", try using the simulation "Concentration" (You will use this information again in Part 5, if your instructor requires it)



Molarity

Cobalt (II) nitrate	5.640 M	Potassium chromate	3.350 M
Cobalt chloride	4.330 M	Nickel (II) chloride	5.210 M
Potassium dichromate	0.510 M	Copper sulfate	1.380 M
Gold (III) chloride	2.30 M	Potassium permanganate	0.480 M

Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises

Yijin Wang



Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises:

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Table of Contents Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises

1. Understanding the eBook Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
 - The Rise of Digital Reading Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
 - Advantages of eBooks Over Traditional Books
2. Identifying Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
 - User-Friendly Interface
4. Exploring eBook Recommendations from Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises

- Personalized Recommendations
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises User Reviews and Ratings
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises and Bestseller Lists
5. Accessing Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Free and Paid eBooks
- Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Public Domain eBooks
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises eBook Subscription Services
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Budget-Friendly Options
6. Navigating Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises eBook Formats
- ePub, PDF, MOBI, and More
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Compatibility with Devices
 - Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
 - Highlighting and Note-Taking Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
 - Interactive Elements Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
8. Staying Engaged with Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
9. Balancing eBooks and Physical Books Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
 - Setting Reading Goals Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
 - Fact-Checking eBook Content of Conclusion Questions And Calculations Concentration And Molarity Post Lab Exercises
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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