

AP Chemistry Final Study Guide

Unit 1: Atomic Structure and Properties (7%–9%)

- At STP, one mole of gas always occupies 22.4 L
- **Mass Spectroscopy:**
 - Atoms have different isotopes (carbon-12, carbon-14, etc) and this will affect the average atomic mass
 - The most abundant isotope will have the greatest effect
- **Coulomb's Law:**
 - The greater the charge of the nucleus, the more energy is required to ionize an electron
 - The closer an electron is to the nucleus, the more energy is required to ionize an electron
- **Electromagnetic Radiation:**
 - When atoms absorb energy, electrons jump to higher energy levels
 - When electrons drop from higher to lower energy levels, atoms give off electromagnetic radiation energy
- **Kinetic Energy of Ionized Electrons:**
 - The faster an ejected electron is going, the more kinetic energy it has
 - Electrons originally further away from the nucleus require less energy to eject, so it will have more kinetic energy
- **Electron Configuration:**
 - When entering the d-block (transition metals), the energy level drops by 1 unless you are using shorthand notation
 - Nickel: $1s^2 2s^2 2p^6 3s^2 3p^6 3d^8 4s^2$ (notice the d-block)
 - Nickel (with shorthand): $[\text{Ar}] 4s^2 3d^8$
- **Law of Definite Proportions:**
 - A pure chemical compound broken down into elements always contains elements of a fixed ratio, independent of where and how it was created
- **Law of Multiple Proportions:**
 - When elements form compounds, the proportions of the elements in those chemical compounds can be expressed in small whole number ratios
- **Aufbau Principle:**
 - When building up the electron configuration of an atom, electrons are placed in orbitals, subshells, and shells in order of increasing energy

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Unveiling the Power of Verbal Art: An Mental Sojourn through **Chemistry Study Guide For Final Exam**

In a world inundated with screens and the cacophony of immediate interaction, the profound power and psychological resonance of verbal art usually diminish in to obscurity, eclipsed by the constant barrage of sound and distractions. However, located within the lyrical pages of **Chemistry Study Guide For Final Exam**, a interesting function of literary elegance that pulses with raw thoughts, lies an memorable journey waiting to be embarked upon. Penned with a virtuoso wordsmith, this magical opus courses visitors on a mental odyssey, delicately exposing the latent possible and profound affect stuck within the complex internet of language. Within the heart-wrenching expanse of this evocative evaluation, we can embark upon an introspective exploration of the book is key themes, dissect their fascinating writing fashion, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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