

Chapters 1 and 2 Study Guide

- Solve the following equation for V_2 : $P_1V_1 = P_2V_2 = \frac{P_1V_1}{P_2}$
- Solve the following equation for m : $\text{heat} = m \times \Delta T \times SH$
 $m = \frac{\text{heat}}{\Delta T \times SH}$
- Write each of the following in scientific notation:
 - 3500
 3.5×10^3
 - 0.000 016
 1.6×10^{-5}
- Write each of the following in scientific notation:
 - 425 000
 4.25×10^5
 - 0.000 000 86
 8.6×10^{-7}
- State the type of measurement indicated in each of the following:
 - 45.6 g **Mass**
 - 1.85 L **Volume**
 - 14 cm **Length**
 - 6.3×10^2 mg **Mass**
- State the type of measurement indicated in each of the following:
 - 15 °C **Temperature**
 - 25 s **Time**
- Identify the significant zeros in each of the following measured numbers:
 - 0.000 250 m
 3 significant zeros
 - 70.040 g
 3 significant zeros
 - 1 000 000 L
 2 significant zeros
- Identify the significant zeros in each of the following measured numbers:
 - 0.04008 m
 5 significant zeros
 - 6.00×10^3 g
 4 sig (100 doesn't count zeros)
- Round off each of the following numbers to three significant figures:
 - 35.7823 m $\rightarrow$ 35.8 m
 - 0.002 621 7 L
 $\rightarrow$ 0.00262 L
 - 3.8268×10^3 g
 $\rightarrow$ 3830 g (3.83 $\times$ 10³) $\rightarrow$ 3830
- Perform the following calculations of measured numbers and give the answers with the correct number of significant figures:
 - $45.26 \times 0.010 88$
 $= 0.49241288$
 0.4924
 - $2.6 + 324 + 0.007 001$
 $= 0.0070$
 - $\frac{4.0 \times 8.00}{16} = 2.0$
 (2 significant zeros)

Chemistry Ch 21 Answers Study Guide

Neil E. Schore



Chemistry Ch 21 Answers Study Guide:

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