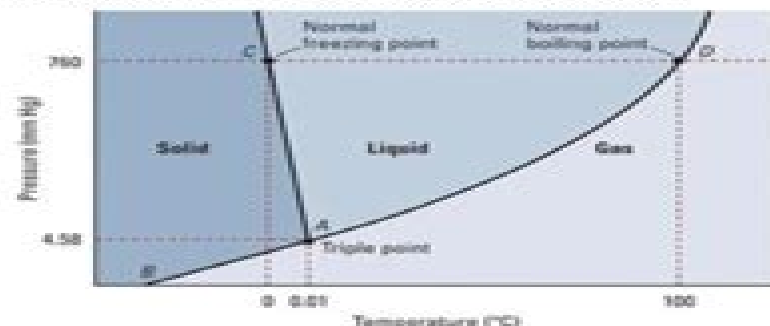


Assessment #3: Phase Diagrams

Multiple Choice: Choose the letter of the correct answer.

For nos. 1,2 and 3, Base your answers on the diagram below.



- b 1. Water exists as a _____ at 700 mmHg and 50 °C.
a. Solid b. liquid c. gas d. supercritical fluid
- a 2. What phase change(s) may occur at pressures below 4.58 mmHg?
a. sublimation/deposition c. melting/freezing
b. all phase changes are possible at these pressures d. vaporization/condensation
- b 3. Based on this phase diagram, which state is the most dense?
a. Gas b. solid c. liquid d. solid and liquid have equal densities

For nos. 4-7 , refer to the diagram below.



- c 4. Based on its phase diagram, which is the most dense state for CO2?
a. Gas b. liquid c. solid d. supercritical fluid
- a 5. What is true regarding CO2 for temperatures above 31 °C?
a. It can never be liquified c. It decomposes.
b. It has a high pressure d. It is a plasma.
- d 6. At 10 atm, dry ice is heated from -100 °C to 30 °C. What changes occur?
a. freezing, then condensation c. melting only
b. sublimation only d. melting, then boiling
- d 7. What is the normal boiling point of CO2?

Chemistry Ws Phase Diagrams Answers

John Hudson Tiner

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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