

CHAPTER 16 KINETICS: RATES AND MECHANISMS OF CHEMICAL REACTIONS

FOLLOW-UP PROBLEMS

16.1A **Plan:** Balance the equation. The rate in terms of the change in concentration with time for each substance is expressed as $-\frac{1}{a} \frac{\Delta[A]}{\Delta t}$, where a is the coefficient of the reactant or product A. The rate of the reactants is given a negative sign.

Solution:

a) The balanced equation is $4\text{NO}(g) + \text{O}_2(g) \rightarrow 2\text{N}_2\text{O}_3$.

Choose O_2 as the reference because its coefficient is 1. Four molecules of NO (nitrogen monoxide) are consumed for every one O_2 molecule, so the rate of O_2 disappearance is 1/4 the rate of NO decrease. By similar reasoning, the rate of O_2 disappearance is 1/2 the rate of N_2O_3 (dinitrogen trioxide) increase.

$$\text{Rate} = -\frac{1}{4} \frac{\Delta[\text{NO}]}{\Delta t} = -\frac{\Delta[\text{O}_2]}{\Delta t} = +\frac{1}{2} \frac{\Delta[\text{N}_2\text{O}_3]}{\Delta t}$$

b) **Plan:** Because NO is decreasing, its rate of concentration change is negative. Substitute the negative value into the expression and solve for $\Delta[\text{O}_2]/\Delta t$.

Solution:

$$\text{Rate} = -\frac{1}{4}(-1.60 \times 10^{-2} \text{ mol/L} \cdot \text{s}) = -\frac{\Delta[\text{O}_2]}{\Delta t} = 4.00 \times 10^{-3} \text{ mol/L} \cdot \text{s}$$

16.1B **Plan:** Examine the equation that expresses the rate in terms of the change in concentration with time for each substance. The number in the denominator of each fraction is the coefficient for the corresponding substance in the balanced equation. The terms that are negative in the rate equation represent reactants in the balanced equation, while the terms that are positive represent products in the balanced equation. For example, the following term, $\frac{1}{a} \frac{\Delta[A]}{\Delta t}$, describes a product (the term is positive) that has a coefficient of a in the balanced equation. In

part b), use the rate equation to compare the rate of appearance of H_2O with the rate of disappearance of O_2 .

Solution:

a) The balanced equation is $4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO} + 6\text{H}_2\text{O}$.

b) **Plan:** Because H_2O is increasing, its rate of concentration change is positive. Substitute its rate into the expression and solve for $\Delta[\text{O}_2]/\Delta t$.

Solution:

$$\begin{aligned} \text{Rate} &= \frac{1}{6}(2.52 \times 10^{-2} \text{ mol/L} \cdot \text{s}) = -\frac{1}{5} \frac{\Delta[\text{O}_2]}{\Delta t} \\ -\frac{5}{6}(2.52 \times 10^{-2} \text{ mol/L} \cdot \text{s}) &= \frac{\Delta[\text{O}_2]}{\Delta t} = -2.10 \times 10^{-2} \text{ mol/L} \cdot \text{s} \end{aligned}$$

The negative value indicates that $[\text{O}_2]$ is decreasing as the reaction progresses. The rate of reaction is always expressed as a positive number, so $[\text{O}_2]$ is decreasing at a rate of $2.10 \times 10^{-2} \text{ mol/L} \cdot \text{s}$.

16.2A **Plan:** The reaction orders of the reactants are the exponents in the rate law. Add the individual reaction orders to obtain the overall reaction order. Use the rate law to determine how the changes listed in the problem will affect the rate.

Solution:

a) The exponent of $[\text{I}^-]$ is 1, so the reaction is **first order with respect to I^-** . Similarly, the reaction is **first order with respect to BrO_3^-** , and **second order with respect to H^+** . The overall reaction order is $(1 + 1 + 2) = 4$, or **fourth order overall**.

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 - The Rise of Digital Reading Chemistrychapter 16 Solutions Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistrychapter 16 Solutions Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistrychapter 16 Solutions Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistrychapter 16 Solutions Manual
 - Personalized Recommendations
 - Chemistrychapter 16 Solutions Manual User Reviews and Ratings
 - Chemistrychapter 16 Solutions Manual and Bestseller Lists
5. Accessing Chemistrychapter 16 Solutions Manual Free and Paid eBooks
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 - Chemistrychapter 16 Solutions Manual eBook Subscription Services
 - Chemistrychapter 16 Solutions Manual Budget-Friendly Options
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 - Chemistrychapter 16 Solutions Manual Enhanced eBook Features
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 - Highlighting and Note-Taking Chemistrychapter 16 Solutions Manual
 - Interactive Elements Chemistrychapter 16 Solutions Manual
 8. Staying Engaged with Chemistrychapter 16 Solutions Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistrychapter 16 Solutions Manual
 9. Balancing eBooks and Physical Books Chemistrychapter 16 Solutions Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistrychapter 16 Solutions Manual
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistrychapter 16 Solutions Manual
 - Setting Reading Goals Chemistrychapter 16 Solutions Manual
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistrychapter 16 Solutions Manual
 - Fact-Checking eBook Content of Chemistrychapter 16 Solutions Manual
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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