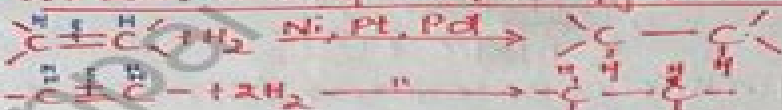


HYDROCARBON

MOP of Alkane

1. From unsaturated hydrocarbon



DM $\rightarrow$ π bond ko H.H se tod do.

exothermic # ROR: Alkyne > Alkene

4 MCTS formed # ROR $\propto \frac{1}{s.H}$

Syn addition $\propto$ stability

Addⁿ of conj. diene



$\text{RaNi} \rightarrow$ alloy of Ni & Al in NaOH
Galvanom/pmp better than Ni(at 100°C)

If RaNi used $\rightarrow$ Sabatier senderson's

$[\text{Rh}(\text{PPh}_3)_3\text{Cl}]$ used also

(Wilkinson's catalyst)

Alkene	Addition	Product
Cis	Syn	Erythro
Trans	Syn	Threo
Cis	Anti	Threo
Trans	Anti	Erythro

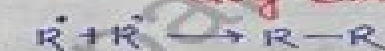
2. From alkyl halide



X hatoo H lagao.

ROR: $\text{R-Cl} < \text{R-Br} < \text{R-I}$ ROR $\propto \frac{1}{BE}$

3. Wurtz reaction



produce higher alkane with even no. of 'C'

$\text{R-X} \rightarrow$ coupling

(1,2)

$\text{R-X} \rightarrow$ disproportionation

(3)

in 2 molecular wurtz faster than in 2 molecular but prod should be stable i.e A.A nahi banana hai.

ROR $\propto \frac{1}{s.H}$ $\text{R-X} > 2^\circ \text{R-X} > 3^\circ \text{R-X}$

for same R: $\text{RF} < \text{RCl} < \text{RBr} < \text{RI}$

4. Decarboxylation reaction



-COOH hatoo H lagao



-COONa hatoo H lagao

ROR $\propto$ stability of intermediate ($\text{R}^\cdot$)

5. Kolbe's electrolysis



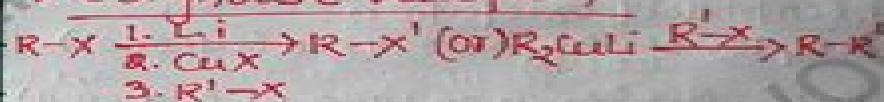
Cathode $\rightarrow \text{H}_2$ Anode $\rightarrow$ Alkane & CO_2

Addition Reaction

6. Frankland reaction



7. Coreyhouse reaction



Attack of $\text{R}^\cdot \rightarrow \text{RDS}$

ROR $\propto \frac{1}{s.H \text{ of } \text{R}'-\text{X}}$

If same R:

$\text{R-F} < \text{R-Cl} < \text{R-Br} < \text{R-I}$

$\text{BE} \downarrow \rightarrow$ leaving tendency $\uparrow$

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**Denise Guinn, Rebecca Brewer, Rachel
C. Lum**



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