

Density and Specific Gravity

3

A. Density of Liquids

A.1 Density of Water

1. Volume (mL)
2. Mass of beaker
3. Mass of beaker + liquid
4. Mass of liquid
5. Density

(Show calculations)

$$\begin{array}{r}
 20.0 \text{ mL} \\
 0.00 \text{ g} \\
 15.82 \text{ g} \\
 \hline
 20.420 \text{ g} \\
 D = \frac{m}{V} = \frac{20.420}{20.0} = 1.01 \text{ g/mL}
 \end{array}$$

A.2 Density of an Unknown Liquid

1. Type of liquid sample
Volume of liquid
2. Mass of beaker
3. Mass of beaker + liquid
4. Mass of liquid
5. Density

(Show calculations)

$$\begin{array}{r}
 \text{Salt water} \\
 20.1 \text{ mL} \\
 0.00 \text{ g} \\
 15.88 \text{ g} \\
 \hline
 21.24 \text{ g} \\
 \frac{21.24 \text{ g}}{20.1 \text{ mL}} = 1.05 \text{ g/mL}
 \end{array}$$

B. Specific Gravity

1. Specific gravity of water
2. Specific gravity of liquid sample
3. Specific gravity of water using a hydrometer
4. Specific gravity of liquid using a hydrometer

$$\begin{array}{r}
 1.01 \\
 1.05 \\
 0.995 \\
 0.790
 \end{array}$$

How do the *calculated* specific gravity of water and the liquid sample (1, 2) compare to the hydrometer readings (3, 4) for water and the liquid sample?

Density And Specific Gravity Laboratory Manual

Caleb Canby Balderston



Density And Specific Gravity Laboratory Manual:

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