

Unit 1 Chemistry Study Guide

History of the Atom:

Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
- **Dalton's Atomic Theory:**
 - o All matter is composed of extremely small particles called atoms.
 - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
 - o Atoms cannot be subdivided, created, or destroyed.
 - o Atoms of different elements can combine in simple whole number ratios to form chemical compounds.
 - o In chemical reactions, atoms are combined, separated or rearranged.
- **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

Billiard Ball Model:



atom

J.J. Thompson (1906):

- **Cathode Ray Experiment –**
 - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

Chemistry Unit Study Guide Answers

G Orfield



Chemistry Unit Study Guide Answers:

The book delves into Chemistry Unit Study Guide Answers. Chemistry Unit Study Guide Answers is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Chemistry Unit Study Guide Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Chemistry Unit Study Guide Answers
 - Chapter 2: Essential Elements of Chemistry Unit Study Guide Answers
 - Chapter 3: Chemistry Unit Study Guide Answers in Everyday Life
 - Chapter 4: Chemistry Unit Study Guide Answers in Specific Contexts
 - Chapter 5: Conclusion
 2. In chapter 1, the author will provide an overview of Chemistry Unit Study Guide Answers. This chapter will explore what Chemistry Unit Study Guide Answers is, why Chemistry Unit Study Guide Answers is vital, and how to effectively learn about Chemistry Unit Study Guide Answers.
 3. In chapter 2, this book will delve into the foundational concepts of Chemistry Unit Study Guide Answers. The second chapter will elucidate the essential principles that need to be understood to grasp Chemistry Unit Study Guide Answers in its entirety.
 4. In chapter 3, this book will examine the practical applications of Chemistry Unit Study Guide Answers in daily life. The third chapter will showcase real-world examples of how Chemistry Unit Study Guide Answers can be effectively utilized in everyday scenarios.
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 6. In chapter 5, this book will draw a conclusion about Chemistry Unit Study Guide Answers. This chapter will summarize the key points that have been discussed throughout the book.
- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Chemistry Unit Study Guide Answers.

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