

Moles & Stoichiometry Cheat Sheet

Calculating Molar Mass 1. Write out formula of compound 2. Determine number of atoms of each element present 3. multiply number of atoms of the element by the atomic mass of the element 4. add up the values just determined for all of the elements present in the compound	Calculating Percent Composition 1. Calculate the molar mass of the compound 2. divide the mass contribution of each element by the total molar mass 3. multiply by 100 to make a percent <i>Note: total of percents should equal 100%!</i>
Mass (grams) $\times$ by molar mass $\leftarrow \text{-----}$ $\text{-----} \rightarrow$ $\div$ by molar mass	Mole $\times$ by Avogadro's # ----- $\leftarrow \text{-----}$ $\text{-----} \rightarrow$ $\div$ by Avogadro's # Particles (atoms, molecules, formula units)
Calculating Empirical Formula 1. Change % of each element to grams of each element (if no sample amount is given assume 100 g) 2. convert grams of each element to moles of each element by dividing the grams by the molar (atomic) mass of that element 3. divide all molar values by the smallest molar value (to establish lowest whole number ratio) 4. if all values are not whole numbers, multiply all by a whole number factor to make the values whole (ex. 1.5 and 1 should be multiplied by 2 to get rid of the .5 = 3 and 2) 5. plug values in with the element they represent	Calculating Molecular Formula 1. calculate the empirical formula first 2. calculate the molar mass of the empirical formula 3. take the molar mass of the molecular formula (given in the problem) and divide it by the molar mass of the empirical formula. This will equal the number of times the e.f. is in the m.f. 4. distribute the number attained through the empirical formula <i>Example: $2(\text{C}_2\text{H}_5) = \text{C}_4\text{H}_{10}$</i>
Calculating the Formula of a Hydrate 1. Determine the mass of water in the hydrate by subtracting the mass of the anhydrous from the mass of the hydrate (if necessary) -OR- convert %'s into grams by assuming a 100g sample 2. convert grams of each into moles of each by dividing by molar mass 3. divide the moles of water by the moles of the anhydrous to determine the number of moles of water in the formula 4. write the formula of the anhydrous followed by a dot then the # of moles of water and "H_2O" <i>Example: $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$</i>	Naming a Hydrate 1. write the name of the ionic compound 2. use a prefix for the number of water molecules (see below) 3. write "hydrate" after the prefix 1 = mono, 2 = di, 3 = tri, 4 = tetra, 5 = penta, 6 = hexa, 7 = hepta, 8 = octa, 9 = nona, 10 = deca <i>Example: Copper (II) sulfate pentahydrate</i>

Cheat Sheet For Stoichiometry

Huangqi Zhang

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

Cheat Sheet For Stoichiometry:

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Table of Contents Cheat Sheet For Stoichiometry

1. Understanding the eBook Cheat Sheet For Stoichiometry
 - The Rise of Digital Reading Cheat Sheet For Stoichiometry
 - Advantages of eBooks Over Traditional Books
2. Identifying Cheat Sheet For Stoichiometry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Cheat Sheet For Stoichiometry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Cheat Sheet For Stoichiometry
 - Personalized Recommendations
 - Cheat Sheet For Stoichiometry User Reviews and Ratings
 - Cheat Sheet For Stoichiometry and Bestseller Lists

5. Accessing Cheat Sheet For Stoichiometry Free and Paid eBooks
 - Cheat Sheet For Stoichiometry Public Domain eBooks
 - Cheat Sheet For Stoichiometry eBook Subscription Services
 - Cheat Sheet For Stoichiometry Budget-Friendly Options
6. Navigating Cheat Sheet For Stoichiometry eBook Formats
 - ePub, PDF, MOBI, and More
 - Cheat Sheet For Stoichiometry Compatibility with Devices
 - Cheat Sheet For Stoichiometry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Cheat Sheet For Stoichiometry
 - Highlighting and Note-Taking Cheat Sheet For Stoichiometry
 - Interactive Elements Cheat Sheet For Stoichiometry
8. Staying Engaged with Cheat Sheet For Stoichiometry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Cheat Sheet For Stoichiometry
9. Balancing eBooks and Physical Books Cheat Sheet For Stoichiometry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Cheat Sheet For Stoichiometry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Cheat Sheet For Stoichiometry
 - Setting Reading Goals Cheat Sheet For Stoichiometry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Cheat Sheet For Stoichiometry
 - Fact-Checking eBook Content of Cheat Sheet For Stoichiometry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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