

CHEMISTRY EQUATIONS & ANSWERS

Chemical equations do not come already balanced. This must be done before the equation can be used in a chemically meaningful way.

All chemical calculations you will see in other units must be done with a balanced equation.

A balanced equation has equal numbers of each type of atom on each side of the equation.

The Law of Conservation of Mass is the rationale for balancing a chemical equation. The law was discovered by Antoine Laurent Lavoisier (1743-94) and this is his formulation of it, translated into English in 1790 from the *Traité élémentaire de Chimie* (which was published in 1789):

"We may lay it down as an incontestable axiom, that, in all the operations of art and nature, nothing is created; an equal quantity of matter exists both before and after the experiment, the quality and quantity of the elements remain precisely the same; and nothing takes place beyond changes and modifications in the combination of these elements." A less wordy way to say it might be:

"Matter is neither created nor destroyed."
Therefore, we must finish our chemical reaction with as many atoms of each element as when we started.

Example #1: Balance the following equation: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

It is an unbalanced equation (sometimes also called a skeleton equation). This means that there are UNEQUAL numbers of at least one atom on each side of the arrow. By the way, a skeleton equation is not wrong, it just hasn't been balanced yet. Presenting it as being balanced would be wrong.

In the example equation, there are two atoms of hydrogen on each side, BUT there are two atoms of oxygen on the left side and only one on the right side.

Remember this: A balanced equation MUST have EQUAL numbers of EACH type of atom on BOTH sides of the arrow.

An equation is balanced by changing coefficients in a somewhat trial-and-error fashion. It is important to note that only the coefficients can be changed, NEVER a subscript.

The coefficient times the subscript gives the total number of atoms.

Three quick examples before balancing the equation.

(a) 2H_2 - there are 2×2 atoms of hydrogen (a total of 4).

(b) $2\text{H}_2\text{O}$ - there are 2×2 atoms of hydrogen (a total of 4) and 2×1 atoms of oxygen (a total of 2).

(c) $2(\text{NH}_4)_2\text{S}$ - there are $2 \times 1 \times 2$ atoms of nitrogen (a total of 4), there are $2 \times 4 \times 2$ atoms of hydrogen (a total of 16), and 2×1 atoms of sulfur (a total of 2).

So, now to balancing the example equation:

$\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ The hydrogen are balanced, but the oxygens are not. We have to get both balanced. We put a two in front of the water and this balances the oxygen.

$\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ However, this causes the hydrogen to become unbalanced. To fix this, we place a two in front of the hydrogen on the left side.

$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ This balances the equation.

Two things you CANNOT do when balancing an equation.

1) You cannot change a subscript.

You cannot change the oxygen's subscript in water from one to two, as in:

$\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$ True, this is a balanced equation, but you have changed the substances in it. H_2O_2 is a completely different substance from H_2O . So, it's not the answer to the question that was asked.

2) You cannot place a coefficient in the middle of a formula.

The coefficient goes at the beginning of a formula, not in the middle, as in:

$\text{H}_2 + \text{O}_2 \rightarrow \text{H}_22\text{O}$ Water only comes as H_2O and you can only use whole formula units of it.

Two more points:

1) Make sure that your final set of coefficients are all whole numbers with no common factors other than one. For example, this equation is balanced: $4\text{H}_2 + 2\text{O}_2 \rightarrow 4\text{H}_2\text{O}$



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