# Chapter 18

#### Study Guide - Answers

- Write the expression for the equilibrium constant for each of the following reactions. 100
  - a.  $H_2(g) + CI_2(g) \cong 2HCl(g)$

$$K_{eq} = \frac{[HCT]^2}{[H_2][CI_2]}$$

b.  $N_2(g) + 3 H_2(g) = 2 NH_1(g)$ 

 $K_{eq} = \frac{[NH_1]^4}{[N_2][H_2]^4}$ c.  $HCN(g) \stackrel{\text{def}}{=} H^4(aq) + CN(aq)$ 

d.  $H_2SO_4(g) \stackrel{\text{def}}{=} H'(aq) + HSO^+(aq)$ 

$$K_{eq} = \frac{-[H'][HSO^4]}{[H;SO_4]}$$

e.  $PbF_2(s) = Pb^2(aq) + 2F(aq)$ 

2 From the data provided below, calculate the value of the equilibrium constant for the reaction.

$$[H_2] = [CI_2] = 1.0 \times 10^{-2}; [HCI] = 1.0 \times 10^{-4}$$

$$K_{eq} = \frac{[HCI]^2}{[H_1][CI_2]}$$

$$K_{eq} = \frac{[1.0 \times 10^{-4}]^2}{[1.0 \times 10^{-2}][1.0 \times 10^{-2}]} = 0.0001$$

b. 
$$N_2(g) + 3 H_2(g) = 2 NH_2(g)$$

$$[N_3] = 4.4 \times 10^{-2}$$
;  $[H_2] = 1.2 \times 10^{-1}$ ;  $[NH_3] = 3.4 \times 10^{-3}$ 

$$K_{eq} = \frac{[3.4 \times 10^{-9}]^2}{[4.4 \times 10^{-9}][1.2 \times 10^{-9}]^2} = 0.152$$

3. For the reaction  $N_2O_4(g) = 2NO_2(g)$ ,  $K_{eq} = 0.2$ . At a particular time, the following concentrations are measured; [N<sub>2</sub>O<sub>4</sub>] = 2.0 M, [NO<sub>2</sub>] = 0.2 M. Is this reaction at equilibrium? If not, in which direction will the reaction proceed?

$$N_2O_2(g) \cong 2NO_2(g)$$

$$[N_2O_3] = 2.0 \text{ M}$$
  
 $[NO_2] = 0.2 \text{ M}$ 

$$Q = -\frac{[NO_2]^2}{[N_2O_4]}$$

$$Q = \frac{[0.2]^2}{[2.0]} = 0.02$$

$$K_{eq} = 0.2$$

 $Q \le K_{eq}$  so the reaction will proceed to the right (forward).

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