

Question 1

1(a)(i) Crude oil and natural gas.

1(a)(ii) Methane.

1(b)(i)

Mass of conical flask and water (g)	326.0
Mass of conical flask (g)	125.0
Mass of water used (g)	201.0
Final temperature of water (°C)	39
Initial temperature of water (°C)	27
Temperature change (°C)	12
Initial mass of candle and watch glass (g)	97.5
Final mass of candle and watch glass (g)	96.0
Mass of candle burnt (g)	1.5

1(b)(ii) Heat absorbed by conical flask = $(125.0 \text{ g}) \times (0.861 \text{ J g}^{-1} \text{ °C}^{-1}) \times (12 \text{ °C}) = 1291.5 \text{ J}$.

1(b)(iii) Heat absorbed by water in the conical flask = $(201.0 \text{ g}) \times (4.2 \text{ J g}^{-1} \text{ °C}^{-1}) \times (12 \text{ °C}) = 10130.4 \text{ J}$.

1(b)(iv) Total heat absorbed by the calorimeter = $1291.5 \text{ J} + 10130.4 \text{ J} = 11421.9 \text{ J}$.

1(b)(v) Heat of combustion = $(11421.9 \text{ J}/1.5 \text{ g}) = 7614.6 \text{ J g}^{-1}$.

1(b)(vi) To reduce heat losses to the surroundings/environment.

- 1(c)(i)
- NO_2 gas evolved i.e., NO_3^- ion present.
 - O_2 gas evolved.
 - CO_2 gas evolved i.e., CO_3^{2-} ion present.

1(c)(ii) • $\text{Fe}^{2+}_{(\text{aq})} + 2\text{OH}^{-}_{(\text{aq})} \rightarrow \text{Fe}(\text{OH})_{2(\text{s})}$.

- 1(c)(iii)
- White precipitate produced.
 - Precipitate darkens/becomes purple (purplish grey) in colour.
 - Precipitate dissolves to give colourless solution.

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Index Medicus ,2001

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