

Chemistry Study Guide

Periodic Table of the Elements

1 H	2 He																	18 Ar	19 K	20 Ca											36 Kr	37 Rb	38 Sr											54 Xe	55 Cs	56 Ba															86 Rn	87 Fr	88 Ra											118 Og																																			
3 Li	4 Be																	12 Mg	13 Al	14 Si	15 P	16 S	17 Cl	18 Ar	19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr	37 Rb	38 Sr	39 Y	40 Zr	41 Nb	42 Mo	43 Tc	44 Ru	45 Rh	46 Pd	47 Ag	48 Cd	49 In	50 Sn	51 Sb	52 Te	53 I	54 Xe	55 Cs	56 Ba	57 La	58 Ce	59 Pr	60 Nd	61 Pm	62 Sm	63 Eu	64 Gd	65 Tb	66 Dy	67 Ho	68 Er	69 Tm	70 Yb	71 Lu	72 Hf	73 Ta	74 W	75 Re	76 Os	77 Ir	78 Pt	79 Au	80 Hg	81 Tl	82 Pb	83 Bi	84 Po	85 At	86 Rn	87 Fr	88 Ra	89 Ac	90 Th	91 Pa	92 U	93 Np	94 Pu	95 Am	96 Cm	97 Bk	98 Cf	99 Es	100 Fm	101 Md	102 No	103 Lr

Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

Chemistry Study Guide Section 2 Answers

Jill Meryl Levy



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Table of Contents Chemistry Study Guide Section 2 Answers

1. Understanding the eBook Chemistry Study Guide Section 2 Answers
 - The Rise of Digital Reading Chemistry Study Guide Section 2 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide Section 2 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide Section 2 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide Section 2 Answers
 - Personalized Recommendations
 - Chemistry Study Guide Section 2 Answers User Reviews and Ratings
 - Chemistry Study Guide Section 2 Answers and Bestseller Lists
5. Accessing Chemistry Study Guide Section 2 Answers Free and Paid eBooks
 - Chemistry Study Guide Section 2 Answers Public Domain eBooks
 - Chemistry Study Guide Section 2 Answers eBook Subscription Services
 - Chemistry Study Guide Section 2 Answers Budget-Friendly Options

6. Navigating Chemistry Study Guide Section 2 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Study Guide Section 2 Answers Compatibility with Devices
 - Chemistry Study Guide Section 2 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Study Guide Section 2 Answers
 - Highlighting and Note-Taking Chemistry Study Guide Section 2 Answers
 - Interactive Elements Chemistry Study Guide Section 2 Answers
8. Staying Engaged with Chemistry Study Guide Section 2 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide Section 2 Answers
9. Balancing eBooks and Physical Books Chemistry Study Guide Section 2 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide Section 2 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Study Guide Section 2 Answers
 - Setting Reading Goals Chemistry Study Guide Section 2 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Study Guide Section 2 Answers
 - Fact-Checking eBook Content of Chemistry Study Guide Section 2 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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