

The forces that hold matter together are called chemical bonds. There are four major types of bonds. We need to learn in detail about these bonds and how they influence the properties of matter. The four major types of bonds are:

- I. Ionic Bonds
- II. Covalent Bonds

- III. Metallic Bonds
- IV. Intermolecular (van der Waals) forces

Ionic Bonds

The ionic bond is formed by the attraction between oppositely charged ions. Ionic bonds are formed between metals and nonmetals. Remember that metal atoms lose one or more valence electrons in order to achieve a stable electron arrangement. When a metal atom loses electrons it forms a positive ion or **cation**. When nonmetals react they gain one or more electrons to reach a stable electron arrangement. When a nonmetal atom gains one or more electrons it forms a negative ion or **anion**. The metal cations donate electrons to the nonmetal anions so they stick together in an ionic compound. This means that ionic bonds are formed by the complete transfer of one or more electrons.

A structure with its particles arranged in a regular repeating pattern is called a **crystal**. Because opposite charges attract and like charges repel, the ions in an ionic compound stack up in a regular repeating pattern called a crystal lattice. The positive ions are pushed away from other positive ions and attracted to negative ions so this produces a regular arrangement of particles where each ion is surrounded by ions of the opposite charge. Each ion in the crystal has a strong electrical attraction to its oppositely charged neighbors so the whole crystal holds together as one giant unit. We have no individual molecules in ionic compounds, just the regular stacking of positive and negative ions.

1. Define the following terms:

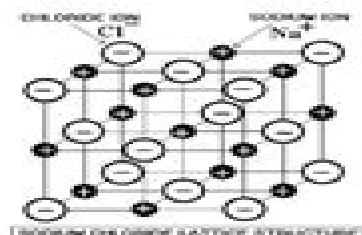
a) ionic bond –

b) cation –

c) anion –

d) crystal –

2. What are the smallest units of an ionic bond?



At room temperature ionic compounds are high melting point solids. They are usually white except for compounds of the transition metals that may be colored. They are brittle (break easily). They do not conduct electricity as solids, but do conduct electricity when melted or dissolved in water.

3. List several properties of ionic compounds:

4. When can electricity be conducted in an ionic bond?

Chemistrynuclear Packet Answers

Jiyuan Zhang



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Table of Contents Chemistrynuclear Packet Answers

1. Understanding the eBook Chemistrynuclear Packet Answers
 - The Rise of Digital Reading Chemistrynuclear Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistrynuclear Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistrynuclear Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistrynuclear Packet Answers
 - Personalized Recommendations
 - Chemistrynuclear Packet Answers User Reviews and Ratings
 - Chemistrynuclear Packet Answers and Bestseller Lists
5. Accessing Chemistrynuclear Packet Answers Free and Paid eBooks
 - Chemistrynuclear Packet Answers Public Domain eBooks
 - Chemistrynuclear Packet Answers eBook Subscription Services
 - Chemistrynuclear Packet Answers Budget-Friendly Options
6. Navigating Chemistrynuclear Packet Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Chemistrynuclear Packet Answers Compatibility with Devices
 - Chemistrynuclear Packet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistrynuclear Packet Answers
 - Highlighting and Note-Taking Chemistrynuclear Packet Answers
 - Interactive Elements Chemistrynuclear Packet Answers
 8. Staying Engaged with Chemistrynuclear Packet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistrynuclear Packet Answers
 9. Balancing eBooks and Physical Books Chemistrynuclear Packet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistrynuclear Packet Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistrynuclear Packet Answers
 - Setting Reading Goals Chemistrynuclear Packet Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistrynuclear Packet Answers
 - Fact-Checking eBook Content of Chemistrynuclear Packet Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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