#### CHAPTER 17

#### PROPERTIES OF SOLUTIONS

#### Solution Composition

12. Mass percent: the percent by mass of the solute in the solution.

> Mole fraction: the ratio of the number of moles of a given component to the total number of modes, of solution

Molarity: the number of moles of solute per liter of solution.

Molality: the number of moles of solute per kilogram of solvent.

Volume is temperature-dependent, whereas mass and the number of moles are not. Only molarity has a volume term, so only molarity is temperature-dependent.

a. 
$$HNO_3(1) \rightarrow H^*(aq) + NO_3^*(aq)$$
 b.  $Na_2SO_4(s) \rightarrow 2 Na^*(aq) + SO_4^{2^*}(aq)$ 

c. 
$$Al(NO_3)_3(s) \rightarrow Al^{3*}(aq) + 3 NO_3^{-}(aq)$$
 d.  $SrBr_2(s) \rightarrow Sr^{2*}(aq) + 2 Br^{-}(aq)$ 

d. 
$$SrBr_2(s) \rightarrow Se^{2^{s}}(aq) + 2 Br'(aq)$$

f. 
$$NH_aBr(s) \rightarrow NH_a^-(aq) + Br^-(aq)$$

g. 
$$NH_4NO_3(s) \rightarrow NH_4'(aq) + NO_3'(aq)$$

h. 
$$CuSO_4(s) \rightarrow Cu^{2^+}(aq) + SO_4^{2^-}(aq)$$

14. 
$$Mol Na_2CO_5 = 0.0700 L \times \frac{3.0 \text{ mol Na}_2CO_3}{L} = 0.21 \text{ mol Na}_2CO_5$$

$$Na_2CO_3(s) \rightarrow 2 Na^*(aq) + CO_3^{2-}(aq); \text{ mol } Na^* = 2(0.21) = 0.42 \text{ mol}$$

$$Mol\ NaHCO_3 = 0.0300\ L \times \frac{1.0\ mol\ NaHCO_3}{L} = 0.030\ mol\ NaHCO_3$$

$$NaHCO_3(s) \rightarrow Na'(aq) + HCO_3'(aq); mol Na' = 0.030 mol$$

$$M_{\mathrm{Na}^+} = \frac{\mathrm{total\ mol\ Na}^+}{\mathrm{total\ volume}} = \frac{0.42\ \mathrm{mol} + 0.030\ \mathrm{mol}}{0.0700\ \mathrm{L} + 0.030\ \mathrm{L}} = \frac{0.45\ \mathrm{mol}}{0.1000\ \mathrm{L}} - 4.5\ M\ \mathrm{Na}^+$$

# **Chapter 17 Properties Of Solutions**

Gilbert Newton Lewis, Merle
Randall, Kenneth S. Pitzer, Leo Brewer

#### **Chapter 17 Properties Of Solutions:**

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