

Finishing Part A
Use the chart below to determine the pH of each solution. Each pH indicator changes color based on the pH of the solution. Look at Methyl Red as an example. Methyl red will be red at and below a pH of 4.8. Additionally, methyl red will be yellow at or above a pH of 6.0. In between a pH of 4.8 and 6.0, methyl red will be orange (red and yellow make orange).

Each pH indicator has a limited range. Use the combination of the indicators to narrow down the pH range for each solution. For example, consider a solution with unknown pH. Start with the indicators that change color around a pH of 7.0, which will be bromothymol blue and phenol red. Our bromothymol blue indicator is blue, so the pH must be greater than 7.6. Likewise, the phenol red indicator is red, so the pH must be greater than 8.0. However, our phenolphthalein solution is colorless, indicating a pH lower than 8.2. Therefore, we have determined that the solution must have a pH between 8.0 and 8.2.

Keywords: child sexual abuse; disclosure; social support

1. Record the experimental pH range for each solution below

Distilled Water	
Tap Water	
0.1M NaOH	
0.1M NH ₄ OH	
0.1M HAc	
0.1M HCl	
0.1M H ₂ SO ₄	

2. Calculate the theoretical pH for each solution and record it below.
- Distilled Water

Distilled Water _____
Tap Water _____
0.1M NaOH _____
0.1M NH_4OH _____
0.1M HAc _____
0.1M HCl _____
0.1M H_2SO_4 _____

Chem Lab 1212 Lab Manual

Georgia College & State University



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CHEM 1212 LAB, KSU, Huck Benjamin Huck,N. J. Kirk,2014-01-01 **Chemistry 1212, LAB Manual** Jack Duff,Alan Gabrielli,2014-08-01 **Principles of Chemistry II** Georgia College & State University,2025-08-15
Lab Manual for Chemistry 1212K Principles of Chemistry II *Manual of Clinical Laboratory Immunology* Noel R. Rose,1997 Reflects changes being thrust upon the laboratory community Chemical Abstracts ,1918 *Barile's Clinical Toxicology* Frank A. Barile,2019-04-24 As with the two previous editions Barile s Clinical Toxicology Principles and Mechanisms Third edition examines the complex interactions associated with clinical toxicological events as a result of therapeutic drug administration or chemical exposure With special emphasis placed on signs and symptoms of diseases and pathology caused by toxins and clinical drugs the new edition examines the complex interactions associated with clinical toxicological events as a result of therapeutic drug administration or chemical exposure The new edition presents the latest up to date protocols for managing various toxic ingestions and the antidotes and treatments associated with their pathology In addition the effect of toxins on a limited number of body systems and drug induced adverse drug reactions are also covered
KEY FEATURES Discusses source of the drug or chemical pharmacological and toxicological mechanisms of action detection identification and treatment Examines the complex interactions associated with clinical toxicological events Emphasizes the signs and symptoms of diseases and pathology caused by toxins and clinical drugs Covers effect of toxins on body systems and drug induced adverse reactions Offers a unique perspective for toxicology pharmacology pharmacy and health professions students The target audience for this book is undergraduate and graduate toxicology students clinical pharmacy Pharm D students emergency medical personnel regulatory agencies and other related health science professionals It satisfies an essential need for a concise yet detailed authoritative fundamental text addressing the current principles of clinical toxicology **Lab World** ,1981 **Accessions of Unlimited Distribution Reports** ,1973-01-19
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