

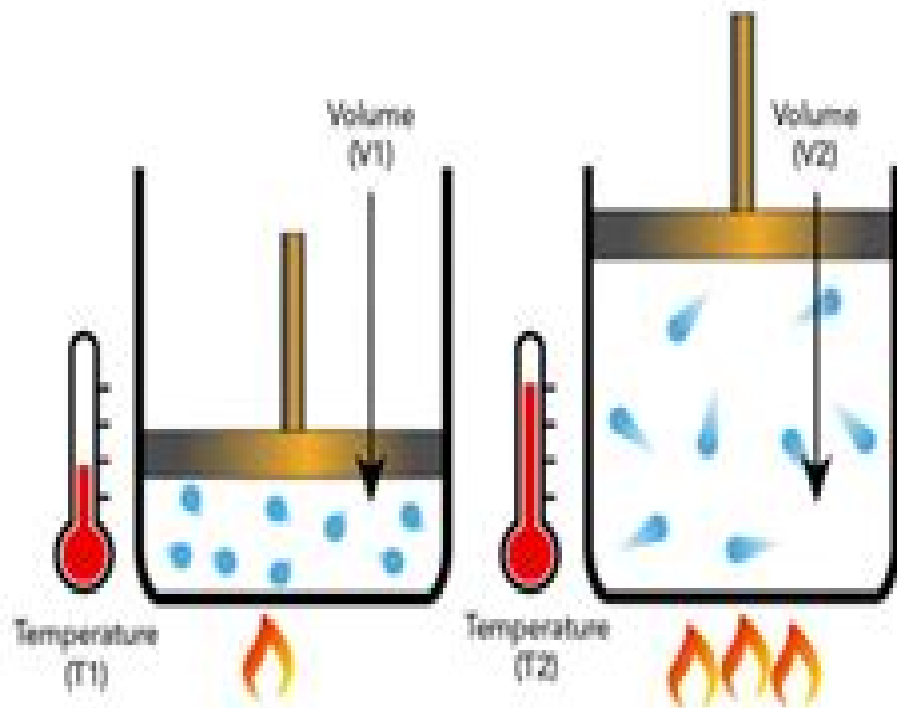
THE CHARLES'S LAW

Charles's Law, named after French scientist Jacques Charles, describes the relationship between the volume and temperature of a gas at constant pressure.

It states that the volume of a gas is directly proportional to its absolute temperature when measured in Kelvin. Mathematically, it is expressed as;

$$V_1/T_1 = V_2/T_2$$

where V_1 and T_1 are the initial volume and temperature, and V_2 and T_2 are the final volume and temperature. Charles's Law is a fundamental principle in gas physics and helps explain the behavior of gases under changing temperature conditions.



Charles Law Chemistry If8766

H. Anthony Neidig, James N. Spencer



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Charles's Law H. Anthony Neidig, James N. Spencer, 1990-01-01 **Gas Laws** Source Wikipedia, 2013-09 Please note that the content of this book primarily consists of articles available from Wikipedia or other free sources online Pages 24 Chapters Acentric factor Amagat's law Avogadro's law Boyle's law Charles's law Combined gas law Compressibility factor Dalton's law Gay Lussac's law Graham's law Henry's law Magic number chemistry Partial pressure Psychrometric constant Redlich Kwong equation of state Van der Waals constants data page Van der Waals equation Excerpt The van der Waals equation is an equation of state for a fluid composed of particles that have a non zero volume and a pairwise attractive inter particle force such as the van der Waals force It was derived in 1873 by Johannes Diderik van der Waals who received the Nobel prize in 1910 for his work on the equation of state for gases and liquids The equation is based on a modification of the ideal gas law and approximates the behavior of real fluids taking into account the nonzero size of molecules and the attraction between them The van der Waals isotherms the model correctly predicts a mostly incompressible liquid phase but the oscillations in the phase transition zone do not fit experimental data The equation uses the following state variables the pressure of the fluid p total volume of the container containing the fluid V number of moles n and absolute temperature of the system T One form of the equation is where v is the volume of the container shared between each particle not the velocity of a particle N is the total number of particles and k_B is Boltzmann's constant given by the universal gas constant R and Avogadro's constant N_A Extra parameters are introduced a is a measure for the attraction between the particles and b is the average volume excluded from v by a particle The equation can be cast into the better known form where a is a measure of the attraction between the particles is the volume excluded by a mole of particles A careful distinction *Charles' Gas Law : Video Demonstration of Experiment* Lydia Anderson, 2013 **Teachers.Net Lesson Exchange: Inquiry Into Charles's Law** , Erica Posthuma presents an inquiry lesson for high school students on Charles's Law which deals with volume temperature and pressure Jacques Alexandre Cesar Charles 1746 1823 was a French physicist chemist and inventor Teachers Net provides the procedures for the lesson as part of the Teachers Net Lesson Exchange online resource

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