

Calculating pH

How will it be used for the concentration differences here?

Why?

As buffers and other substances control pH in a system, it is a way to regulate the acidity or basicity of a cell. Also, this property is also controlled through pH levels in sports as athletes regulate their acid-base balance after an extended pH value that, when it becomes being consumed. For example, when a pH of 7 is reached, before it could reach lower, it is kept at a pH of 7.5 for the concentration of the buffer to be maintained in the system.

Table 1 – Ion Concentrations for Acids and Bases

Acids	Substance	Acidic, Basic or Neutral?	$[H^+]$ (M)	$[OH^-]$ (M)	$[H^+] > [OH^-]$ or $[OH^-] > [H^+]$
1	0.1 M HCl (strong)	Acidic	$1.0 \times 10^{-1} M$	$1.0 \times 10^{-13} M$	$[H^+] > [OH^-]$
2	0.001 M HCl (strong)	Acidic	$1.0 \times 10^{-3} M$	$1.0 \times 10^{-11} M$	$[H^+] > [OH^-]$
3	0.00001 M HCl (strong)	Acidic	$1.0 \times 10^{-5} M$	$1.0 \times 10^{-9} M$	$[H^+] > [OH^-]$
4	0.000001 M HCl (strong)	Acidic	$1.0 \times 10^{-6} M$	$1.0 \times 10^{-8} M$	$[H^+] > [OH^-]$
5	0.0000001 M HCl (strong)	Neutral	$1.0 \times 10^{-7} M$	$1.0 \times 10^{-7} M$	$[H^+] = [OH^-]$
6	0.00000001 M HCl (strong)	Basic	$1.0 \times 10^{-8} M$	$1.0 \times 10^{-6} M$	$[OH^-] > [H^+]$
7	0.000000001 M HCl (strong)	Basic	$1.0 \times 10^{-9} M$	$1.0 \times 10^{-5} M$	$[OH^-] > [H^+]$
8	0.0000000001 M HCl (strong)	Basic	$1.0 \times 10^{-10} M$	$1.0 \times 10^{-4} M$	$[OH^-] > [H^+]$
9	0.00000000001 M HCl (strong)	Basic	$1.0 \times 10^{-11} M$	$1.0 \times 10^{-3} M$	$[OH^-] > [H^+]$
10	0.000000000001 M HCl (strong)	Basic	$1.0 \times 10^{-12} M$	$1.0 \times 10^{-2} M$	$[OH^-] > [H^+]$

1. What does the neutral pH of 7 indicate? (Indicate)

It indicates that the solution is neutral.

2. In Solution 1, which one has a higher concentration, hydrochloric acid or hydroxide ions?

Hydrochloric acid.

3. How did you determine the concentration of hydrochloric acid ions in Solution 1 or Solution 2? How did you determine the concentration of the acid?

From the pH value, I determined all of the hydrochloric acid ions in Solution 1 and Solution 2. I used the pH value to determine the concentration of the acid.

4. How did you determine the concentration of hydroxide ions in Solution 1 or Solution 2? How did you determine the concentration of the base?

From the pH value, I determined the concentration of the hydroxide ions in Solution 1 and Solution 2. I used the pH value to determine the concentration of the base.

Calculating Ph Pogil Answers

G. D. Peckham

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Calculating Ph Pogil Answers:

A Generalised Equation for Calculating PH G. D. Peckham, 1984

Ph Hans Skrifvars, 1999

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