

12 • The Gas Laws**THE COMBINED GAS LAW**

In practical terms, it is often difficult to hold any of the variables constant. When there is a change in pressure, volume and temperature, the combined gas law is used.

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \text{or} \quad P_1 \times V_1 \times T_2 = P_2 \times V_2 \times T_1$$

$$K = ^\circ C + 273$$

*See
Scratch
paper*

Complete the following chart.

	P_1	V_1	T_1	P_2	V_2	T_2
1	1.50 atm	3.00 L	20.0°C 293 K	2.50 atm		30.0°C 303 K
2	720. torr	256. mL	25.0°C 298 K		250. mL	50.0°C 323 K
3	600. mmHg	2.50 L	22.0°C 325 K	760. mmHg	1.80 L	
4		750. mL	0.00°C 273 K	2.00 atm	500. mL	25.0°C 298 K
5	95.0 kPa	4.00 L		101. kPa	6.00 L	498.°C 471. K or
6	650. torr		100.°C 373 K	900. torr	225. mL	150.°C 423 K
7	850. mmHg	1.50 L	15.0°C 288 K		2.50 L	30.0°C 303 K
8	125. kPa	125. mL		100. kPa	100 mL	75.0°C 348 K

Combined Gas Law Answer Key

Effiong Eyo

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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Table of Contents Combined Gas Law Answer Key

1. Understanding the eBook Combined Gas Law Answer Key
 - The Rise of Digital Reading Combined Gas Law Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Combined Gas Law Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Combined Gas Law Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Combined Gas Law Answer Key
 - Personalized Recommendations
 - Combined Gas Law Answer Key User Reviews and Ratings
 - Combined Gas Law Answer Key and Bestseller Lists
5. Accessing Combined Gas Law Answer Key Free and Paid eBooks
 - Combined Gas Law Answer Key Public Domain eBooks
 - Combined Gas Law Answer Key eBook Subscription Services
 - Combined Gas Law Answer Key Budget-Friendly Options
6. Navigating Combined Gas Law Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - Combined Gas Law Answer Key Compatibility with Devices
 - Combined Gas Law Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Combined Gas Law Answer Key
 - Highlighting and Note-Taking Combined Gas Law Answer Key
 - Interactive Elements Combined Gas Law Answer Key
 8. Staying Engaged with Combined Gas Law Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Combined Gas Law Answer Key
 9. Balancing eBooks and Physical Books Combined Gas Law Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Combined Gas Law Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Combined Gas Law Answer Key
 - Setting Reading Goals Combined Gas Law Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Combined Gas Law Answer Key
 - Fact-Checking eBook Content of Combined Gas Law Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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