

Calculating pH

How will it be used for the concentration differences here?

Why?

As buffers and other substances control pH in organisms, it may be possible to control the acidity or alkalinity of a cell. How? This property, called **buffered capacity**, pH levels are kept at a relatively constant range after changing either an external pH value, that, when it arrives there, is converted to a constant pH. How does that pH of 7.4 work? Below 7.4 and above 7.4 is fatal for cells. When the concentration of buffers (pH) is greater than the normal level.

Table 1 – Ion Concentrations for Acids and Bases

Acids	Substance	Acidic, Basic or Neutral?	$[H^+]$ (M)	$[OH^-]$ (M)	$[H^+] > [OH^-]$?
1	0.1 M HCl (strong)	Acidic	$1 \times 10^{-1} M$	$1 \times 10^{-13} M$	Yes
2	0.001 M HCl (strong)	Acidic	$1 \times 10^{-3} M$	$1 \times 10^{-11} M$	Yes
3	0.00001 M HCl (strong)	Acidic	$1 \times 10^{-5} M$	$1 \times 10^{-9} M$	Yes
4	0.000001 M HCl (strong)	Acidic	$1 \times 10^{-6} M$	$1 \times 10^{-8} M$	Yes
5	0.0000001 M HCl (strong)	Neutral	$1 \times 10^{-7} M$	$1 \times 10^{-7} M$	Yes
6	0.00000001 M HCl (strong)	Basic	$1 \times 10^{-8} M$	$1 \times 10^{-6} M$	Yes
7	0.000000001 M HCl (strong)	Basic	$1 \times 10^{-9} M$	$1 \times 10^{-5} M$	Yes
8	0.0000000001 M HCl (strong)	Basic	$1 \times 10^{-10} M$	$1 \times 10^{-4} M$	Yes
9	0.00000000001 M HCl (strong)	Basic	$1 \times 10^{-11} M$	$1 \times 10^{-3} M$	Yes
10	0.000000000001 M HCl (strong)	Basic	$1 \times 10^{-12} M$	$1 \times 10^{-2} M$	Yes

1. What does the neutral pH of 7.4 indicate? (Acidity)

Acidity is a measure of the concentration of H^+ ions.

2. In Table 1, which one has a higher concentration of hydroxide ions (OH^-)?

Acidity

3. How did you determine the concentration of hydroxide ions and what did you find? (The pH of 7.4 is neutral, so the concentration of H^+ and OH^- is the same.)

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G. D. Peckham



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A Generalised Equation for Calculating PH G. D. Peckham, 1984

Ph Hans Skrifvars, 1999

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