



Lesson Plan: What have these simulations and the gas laws shown us?

Aim: How are pressure, volume, and temperature related?

Agenda:

- Do Now - 5 minutes
- Pressure: Internal vs. external - 10 minutes
- Gas Laws - Boyle's and Charles's Law - 15 minutes
- Demo: Balloon in hot water - 5 minutes
- Summary: Gases and how they behave - 15 minutes
- Homework: Question on gas laws

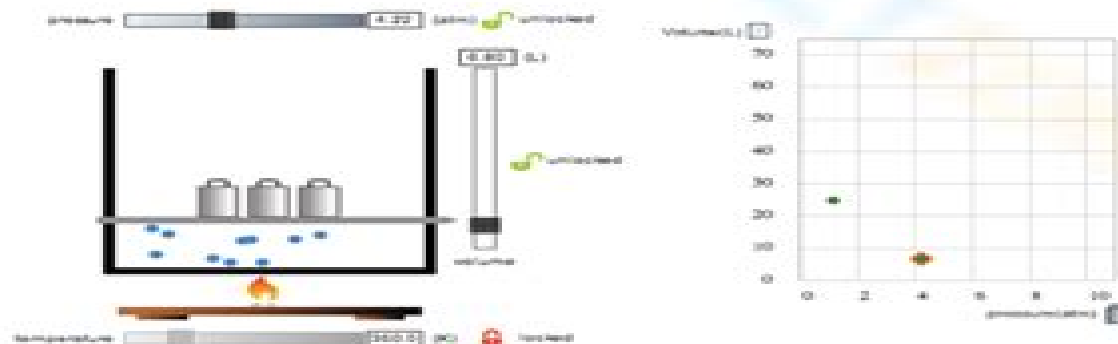
Pressure: Internal vs. external (start boiling water)

Draw a balloon... arrows going in and out. When something isn't moving, like a balloon, the pressure from the air all around us is equal to the amount the gas inside the balloon is pushing out.

So let's say I have a box that has a lid that slides up and down, if I apply pressure on the top and the lid stops sliding, that means that the gas inside the box is putting the same pressure on the inside to hold you there.

So if you apply pressure on the top of the box, do you see why pressure is going up inside the box as well?

If you apply external pressure, why doesn't the volume go to zero? What is holding it up? [no longer conforming to the Ideal Gas Law]



Demonstration Analogy

Use 8 volunteers... put them in a square. (4 inside [they are the gas molecules], 4 make the walls, I push them in) As the wall goes in what happens to the gas molecules? Do they like this?

Charles Law Computer Activity Answers

LL Leslie



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