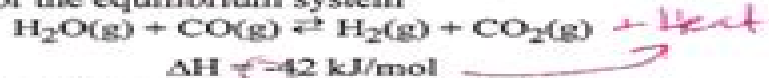


9. $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2 \text{NO}_2(\text{g})$
- At 25°C , 0.11 mole of N_2O_4 reacts to form 0.10 mol of N_2O_4 and 0.02 mole of NO_2 . At 90°C , 0.11 mole of N_2O_4 forms 0.050 mole of N_2O_4 and 0.12 mole of NO_2 . From these data we can conclude
- N_2O_4 molecules react by a second order rate law.
 - N_2O_4 molecules react by a first order rate law.
 - the reaction is exothermic.
 - N_2O_4 molecules react faster at 25°C than at 90°C .
 - the equilibrium constant for the reaction above increases with an increase in temperature.

10. For the equilibrium system



K_c equals 0.62 at 1260 K. If 0.10 mole each of H_2O , CO , H_2 and CO_2 (each at 1260 K) were placed in a 1.0-Liter flask at 1260 K, when the system came to equilibrium...

	The temperature would	The mass of CO would
a)	decrease	increase
b)	decrease	decrease
c)	remain constant	increase
d)	increase	decrease
e)	increase	increase

$$Q = \frac{(0.10)(0.10)}{(0.10)(0.10)} = 1 \quad K_c = 0.62$$

$Q > K_c$ will shift to
 To use up products
 and make reactants
 Temp $\downarrow$ CO $\uparrow$

11. For the reaction system,



the conditions that would favor maximum conversion of the reactants to products would be

- high temperature and high pressure
- high temperature, pressure unimportant
- high temperature and low pressure
- low temperature and high pressure
- low temperature and low pressure

12. Solid HgO , liquid Hg , and gaseous O_2 are placed in a glass bulb and are allowed to reach equilibrium at a given temperature.



The mass of HgO in the bulb could be increased by

- adding more Hg . NO CHANGE
- removing some O_2 .
- reducing the volume of the bulb.
- increasing the temperature.
- removing some Hg . NO CHANGE

Answers: (Please use CAPITAL letters)

1.	E
2.	D
3.	C
4.	B
5.	D
6.	D

7.	D
8.	A
9.	E
10.	A
11.	D
12.	C

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