

Chemistry Study Guide

Periodic Table of the Elements

1 H 1.008 3 Li 6.941 4 Be 9.012 5 B 10.81 6 C 12.01 7 N 14.01 8 O 16.00 9 F 18.998 10 Ne 20.18
11 Na 22.99 12 Mg 24.31 13 Al 26.98 14 Si 28.09 15 P 30.97 16 S 32.07 17 Cl 35.45 18 Ar 39.95
19 K 39.10 20 Ca 40.08 21 Sc 44.96 22 Ti 47.88 23 V 50.94 24 Cr 52.00 25 Mn 54.94 26 Fe 55.85 27 Co 58.93 28 Ni 58.71 29 Cu 63.55 30 Zn 65.39 31 Ga 69.72 32 Ge 72.64 33 As 74.92 34 Se 78.96 35 Br 79.90 36 Kr 83.80
37 Rb 85.47 38 Sr 87.62 39 Y 88.91 40 Zr 91.22 41 Nb 92.91 42 Mo 95.94 43 Tc 98.91 44 Ru 101.07 45 Rh 102.91 46 Pd 106.42 47 Ag 107.87 48 Cd 112.41 49 In 114.82 50 Sn 118.71 51 Sb 121.76 52 Te 127.60 53 I 126.91 54 Xe 131.29
55 Cs 132.91 56 Ba 137.33 57 La 138.91 58 Ce 140.12 59 Pr 140.91 60 Nd 144.24 61 Pm 144.91 62 Sm 150.36 63 Eu 151.96 64 Gd 157.25 65 Tb 158.93 66 Dy 162.50 67 Ho 164.93 68 Er 167.26 69 Tm 168.93 70 Yb 173.05 71 Lu 174.97
72 Fr 223.02 73 Ra 226.03 74 Ac 227.03 75 Th 232.04 76 Pa 231.04 77 U 238.03 78 Np 237.05 79 Pu 244.06 80 Am 243.06 81 Cm 247.07 82 Bk 247.07 83 Cf 251.08 84 Es 252.08 85 Fm 257.09 86 Md 258.10 87 No 259.10 88 Lr 262.11
79 Ir 221.03 80 Pt 231.04 81 Au 231.04 82 Hg 231.04 83 Tl 231.04 84 Pb 231.04 85 Bi 231.04 86 Po 231.04 87 At 231.04 88 Rn 231.04
89 La 138.91 90 Ce 140.12 91 Pr 140.91 92 Nd 144.24 93 Pm 144.91 94 Sm 150.36 95 Eu 151.96 96 Gd 157.25 97 Tb 158.93 98 Dy 162.50 99 Ho 164.93 100 Er 167.26 101 Tm 168.93 102 Yb 173.05 103 Lu 174.97
91 Ac 227.03 92 Th 232.04 93 Pa 231.04 94 U 238.03 95 Np 237.05 96 Pu 244.06 97 Am 243.06 98 Cm 247.07 99 Bk 247.07 100 Cf 251.08 101 Es 252.08 102 Fm 257.09 103 Md 258.10 104 No 259.10 105 Lr 262.11
101 Uut 288.10 102 Fl 289.10 103 Uup 289.10 104 Lv 289.10 105 Uus 289.10 106 Uuo 289.10

Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

Chemistry Study Guide Review Answers

T. W. Graham Solomons



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