

Name _____ answer key _____ Date _____

Stoichiometry: Mass – Mole Worksheet

Based on the following chemical equation, please answer the following questions!



Calculate the amount of moles of NaCl needed to produce 75 g of FeCl₃.

$$\frac{9 \text{ FeCl}_3}{1 \text{ mole FeCl}_3} \times \frac{1 \text{ mole FeCl}_3}{9 \text{ FeCl}_3} \times \frac{\text{mole NaCl}}{\text{mole FeCl}_3}$$

Calculate the amount of Fe₂(SO₄)₃ in grams produce 215 g Na₂SO₄.

$$\frac{9 \text{ Na}_2\text{SO}_4}{1 \text{ mole Na}_2\text{SO}_4} \times \frac{1 \text{ mole Na}_2\text{SO}_4}{9 \text{ Na}_2\text{SO}_4} \times \frac{\text{mole Fe}_2(\text{SO}_4)_3}{\text{mole Na}_2\text{SO}_4} \times \frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3}$$

Calculate the amount in grams of FeCl₃ produced with 95g of Fe (SO₄)₃

$$\frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3} \times \frac{1 \text{ mole Fe}_2(\text{SO}_4)_3}{9 \text{ Fe}_2(\text{SO}_4)_3} \times \frac{\text{mole FeCl}_3}{\text{mole Fe}(\text{SO}_4)_3} \times \frac{9 \text{ FeCl}_3}{1 \text{ mole FeCl}_3}$$

Calculate the amount of moles of Fe₂(SO₄)₃ needed to react with 123.5 g of NaCl

$$\frac{9 \text{ NaCl}}{1 \text{ mole NaCl}} \times \frac{1 \text{ mole NaCl}}{9 \text{ NaCl}} \times \frac{\text{mole Fe}(\text{SO}_4)_3}{\text{mole NaCl}}$$

Based on the following chemical equation, please answer the following questions!



$$\frac{9 \text{ Ca}(\text{HCO}_3)_2}{1 \text{ mole Ca}(\text{HCO}_3)_2} \times \frac{1 \text{ mole Ca}(\text{HCO}_3)_2}{9 \text{ Ca}(\text{HCO}_3)_2} \times \frac{\text{mole CaCO}_3}{\text{mole Ca}(\text{HCO}_3)_2}$$

Chemistry Study Guide The Mole Answers

James E. Brady, Fred Senese



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Table of Contents Chemistry Study Guide The Mole Answers

1. Understanding the eBook Chemistry Study Guide The Mole Answers
 - The Rise of Digital Reading Chemistry Study Guide The Mole Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide The Mole Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Study Guide The Mole Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide The Mole Answers
 - Personalized Recommendations
 - Chemistry Study Guide The Mole Answers User Reviews and Ratings
 - Chemistry Study Guide The Mole Answers and Bestseller Lists
5. Accessing Chemistry Study Guide The Mole Answers Free and Paid eBooks
 - Chemistry Study Guide The Mole Answers Public Domain eBooks
 - Chemistry Study Guide The Mole Answers eBook Subscription Services

- Chemistry Study Guide The Mole Answers Budget-Friendly Options
- 6. Navigating Chemistry Study Guide The Mole Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Study Guide The Mole Answers Compatibility with Devices
 - Chemistry Study Guide The Mole Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Study Guide The Mole Answers
 - Highlighting and Note-Taking Chemistry Study Guide The Mole Answers
 - Interactive Elements Chemistry Study Guide The Mole Answers
- 8. Staying Engaged with Chemistry Study Guide The Mole Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Study Guide The Mole Answers
- 9. Balancing eBooks and Physical Books Chemistry Study Guide The Mole Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Study Guide The Mole Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Study Guide The Mole Answers
 - Setting Reading Goals Chemistry Study Guide The Mole Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Study Guide The Mole Answers
 - Fact-Checking eBook Content of Chemistry Study Guide The Mole Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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