

Test Bank - Chapter 6

The questions in the test bank cover the concepts from the lessons in Chapter 6. Select questions from any of the categories that match the content you covered with students. The types of questions include multiple choice, true/false, fill-in-the-blank, and short answer.

Multiple Choice

1. A chemical change is different than a physical change because in a chemical change
 - a) Chemicals are used
 - b) Molecules do not physically touch
 - c) A new substance is formed and in a physical change no new substance is formed
 - d) The change can be seen but in a physical change it cannot
2. In a chemical reaction
 - a) The atoms of the reactants always stay together to form the products
 - b) The atoms of the reactants unbind, rearrange, and then rebind to form the products
 - c) New atoms are formed which combine to make the products
 - d) Some atoms disappear while others multiply to form the products
3. In the chemical reaction $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$, there are reactants on the left of the equation and products on the right. If you count the atoms in the reactants, there are
 - a) 4 carbon atoms, 4 hydrogen atoms, and 2 oxygen atoms
 - b) 1 carbon atom, 4 hydrogen atoms, and 2 oxygen atoms
 - c) 1 carbon atom, 4 hydrogen atoms, and 4 oxygen atoms
 - d) 4 carbon atoms, 4 hydrogen atoms, and 4 oxygen atoms
4. In a chemical reaction, mass is conserved. This means that
 - a) The mass of the reactants stays the same during a chemical reaction
 - b) The mass of the products stays the same during a chemical reaction
 - c) The type and number of atoms in the reactants equals the type and number of atoms in the products
 - d) The mass of the products is always twice the mass of the reactants.
5. If the reactants on the left side of a chemical equation are $\text{C}_3\text{H}_8 + 5\text{O}_2$, the products in a balanced equation could be
 - a) $4\text{CO}_2 + 3\text{H}_2\text{O}$
 - b) $3\text{CO}_2 + 4\text{H}_2\text{O}$
 - c) $2\text{CO}_2 + 3\text{H}_2\text{O}$
 - d) $3\text{CO} + 4\text{H}_2\text{O}$

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