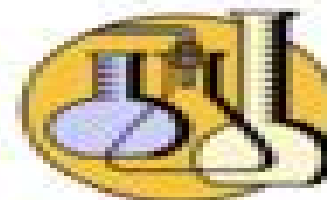


Stoichiometry and The Mole – Review Guide

Remember:

- 1 mole = 6.02×10^{23} atoms, molecules, formula units
- 1 mole = grams (of molar mass)
- 1 mole = 22.4 liters (L) of gas



- 1) The process used to predict the amount of a product that will be formed is stoichiometry.
- 2) If you know the mass of a substance, you can find the number of moles by dividing it by the molar mass of substance.
- 3) If you know the number of particles (atoms) of a substance, you can find the number of moles by dividing it by Avogadro's Number.
- 4) Summarize the steps of a gram-to-gram stoichiometry problem:

Grams A $\rightarrow$ Moles A $\rightarrow$ Moles B $\rightarrow$ Grams B

- 5) The essential stoichiometry problem needed to convert from one chemical to another is:
Mole to Mole
- 6) Two reactants undergo a chemical reaction. At the end of the reaction, Chemical A still exists in the container, while Chemical B has been completely used up. Which is the limiting reactant?
Chemical B
- 7) What is a mole? A unit of measurement used to count atoms/molecules.
- 8) What is Avogadro's Number? 6.02×10^{23}
- 9) What does Avogadro's number tell us? The number of atoms or molecules in a mole.

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