

REVIEW FOR THERMOCHEMISTRY TEST

1. The combustion of methanol is shown by the following equation:



- (a) Using the data below, find the heat of reaction for the combustion of methanol.



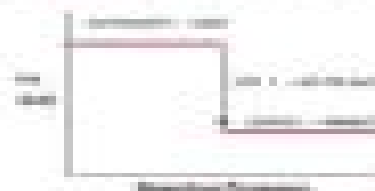
- (b) What is the molar enthalpy of combustion for methanol?

$$\Delta H = -1278 \text{ kJ/mol} \quad \text{or} \quad -639 \text{ kJ/mol}$$

- (c) Express the energy change for the target reaction in the form $\Delta H_{\text{reaction}}$.

$$\Delta H_{\text{reaction}} \text{ CH}_3\text{OH} = -639 \text{ kJ/mol}$$

- (d) Draw the potential energy diagram to illustrate the energy change for the target reaction.



Chemistry Thermochemistry Unit Test Review

M Tight



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